

# Covalent Bonding

## How does a covalent bond form?

- A **covalent bond** forms when two atoms **share** one or more pairs of electrons.
- Atoms in covalent bonds typically have **similar electronegativities** and are **usually non-metals**.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

## Types of Covalent Bonds

- **Single Bond:** One pair of electrons is shared (e.g.,  $\text{H}_2$ ,  $\text{Cl}_2$ ).
- **Double Bond:** Two pairs of electrons are shared (e.g.,  $\text{O}_2$ ,  $\text{CO}_2$ ).
- **Triple Bond:** Three pairs of electrons are shared (e.g.,  $\text{N}_2$ ).



## Dot-and-Cross Diagrams

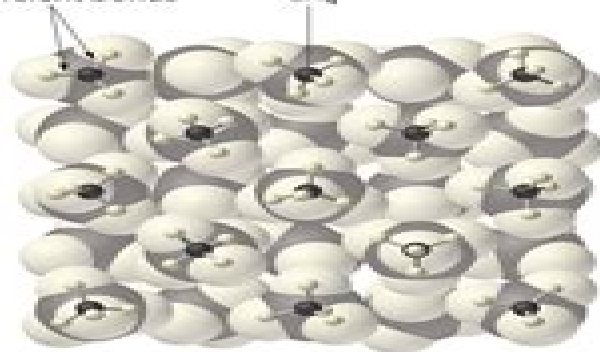
- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

## Example: Water $\text{H}_2\text{O}$



Covalent bonds

$\text{CH}_4$



## Properties of Covalent Compounds

- **Molecular Structure:** They are made up of discrete molecules with **weak intermolecular forces**.
- **Melting and Boiling Points:** Generally **lower** than ionic compounds because the forces between molecules are weaker.
- **Electrical Conductivity:** **Poor conductors** of electricity in all states, as they do not have free ions or electrons.
- **Solubility:** Often **soluble in nonpolar solvents** and may be insoluble in water, depending on their polarity.

# Study Guide Covalent Bonding Key

**William H. Myers, William T. Mooney**



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## **Table of Contents Study Guide Covalent Bonding Key**

1. Understanding the eBook Study Guide Covalent Bonding Key
  - The Rise of Digital Reading Study Guide Covalent Bonding Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Covalent Bonding Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Covalent Bonding Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Covalent Bonding Key
  - Personalized Recommendations
  - Study Guide Covalent Bonding Key User Reviews and Ratings
  - Study Guide Covalent Bonding Key and Bestseller Lists

5. Accessing Study Guide Covalent Bonding Key Free and Paid eBooks
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  - Study Guide Covalent Bonding Key Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
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  - Study Guide Covalent Bonding Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Covalent Bonding Key
  - Highlighting and Note-Taking Study Guide Covalent Bonding Key
  - Interactive Elements Study Guide Covalent Bonding Key
8. Staying Engaged with Study Guide Covalent Bonding Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Covalent Bonding Key
9. Balancing eBooks and Physical Books Study Guide Covalent Bonding Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Covalent Bonding Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Covalent Bonding Key
  - Setting Reading Goals Study Guide Covalent Bonding Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Covalent Bonding Key
  - Fact-Checking eBook Content of Study Guide Covalent Bonding Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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