

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Solutions

## Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

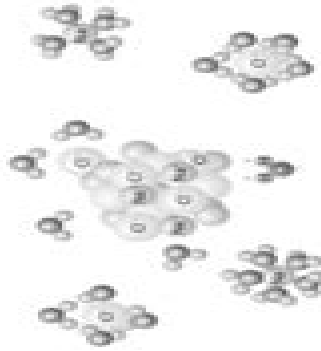
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
miscible	insoluble	solvent	solute

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is hydration in which the solvent is water. What is hydration?  
**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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Chemistry: Matter and Change • Chapter 15

85

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? **The sodium and chloride ions are separated and surrounded by the water molecules.**

water molecules.

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. **Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? **How do you know?**

**The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**

19. List three ways that the rate of solvation may be increased.

**stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

- true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- true** 17. In a saturated solution, solvation and crystallization are in equilibrium.
- true** 18. Additional water can be dissolved in an unsaturated solution.
- true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

86 Chemistry: Matter and Change • Chapter 15

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# Study Guide For Content

**Curtis O. Byer, Louis W. Shainberg**



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## **Table of Contents Study Guide For Content**

1. Understanding the eBook Study Guide For Content
  - The Rise of Digital Reading Study Guide For Content
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For Content
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content
  - Personalized Recommendations
  - Study Guide For Content User Reviews and Ratings
  - Study Guide For Content and Bestseller Lists
5. Accessing Study Guide For Content Free and Paid eBooks
  - Study Guide For Content Public Domain eBooks
  - Study Guide For Content eBook Subscription Services
  - Study Guide For Content Budget-Friendly Options

6. Navigating Study Guide For Content eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide For Content Compatibility with Devices
  - Study Guide For Content Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide For Content
  - Highlighting and Note-Taking Study Guide For Content
  - Interactive Elements Study Guide For Content
8. Staying Engaged with Study Guide For Content
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For Content
9. Balancing eBooks and Physical Books Study Guide For Content
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For Content
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content
  - Setting Reading Goals Study Guide For Content
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide For Content
  - Fact-Checking eBook Content of Study Guide For Content
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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