

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4 Determine moles of product & mass of product.

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$

If you have 100 grams of glucose and 400 grams of oxygen, determine the limiting reactant.

6 How many grams of water (steam) are produced?

7 Determine moles of product & mass of product.



8 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
If you start with 10 moles of magnesium and 2 moles of sodium, determine the limiting reactant.

9 Calculate the grams of magnesium chloride produced.

10 How many grams of the excess reactant is produced in a left over?

11 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
If you have 100 grams of magnesium and 400 grams of sodium, determine the limiting reactant.

12 How much sodium chloride is produced?

13 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
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If you have 100 grams of glucose and 400 grams of oxygen, determine the limiting reactant.

14 Determine moles of product & mass of product.

15 Calculate the grams of water (steam) are produced?

16 How many grams of the excess reactant is produced in a left over?

17 Determine moles of product & mass of product.

18 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure sodium is produced.
 $Mg + 2NaCl \rightarrow MgCl_2 + 2Na$
If you have 100 grams of magnesium and 400 grams of sodium, determine the limiting reactant.

19 Calculate the grams of magnesium chloride produced.

20 How many grams of the excess reactant is produced in a left over?

ANSWER KEY INCLUDED

Study Guide Answer Key Limiting Reactant

John A. Olmsted, Gregory M. Williams



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Chemistry, Student Study Guide** James E. Brady, Fred

Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane CH_4 from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee. Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans. Now that's real chemical understanding. Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist. This text helps you understand how chemical phenomena are governed by what happens at the molecular level, apply critical thinking skills to chemical concepts and problems, and master the basic mathematical techniques needed for quantitative reasoning. You'll see the world as chemists do and learn to appreciate the chemical processes all around us. A Fourth Edition with a lot of new perks. Revisions include a new early energy chapter, revised coverage of bonding, expanded coverage of intermolecular forces, and increased coverage of multiple equilibria, including polyprotic acids. New pedagogy strengthens students' critical thinking and problem-solving skills.

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Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter,Steven S. Zumdahl,2004-04

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new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Study Guide General Chemistry, Second Edition, Becker/Wentworth* Becker, Benjamin L. Carroll, 1980

Prentice Hall Chemistry, 2000 Understanding Chemistry Fred M. Dewey, 1994 **Essentials of Introductory Chemistry** Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry **Foundations of College Chemistry, Alternate** Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis **Foundations of Physical Science** Tom Hsu PH.D., 2002 **Holt Chemistry** Ralph Thomas Myers, 2004

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