

Ex 1. For which of the following reactions would the ΔH° and the equilibrium constant K_{eq} be

1) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} > 1$ \rightarrow **spontaneous**

2) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} < 1$ \rightarrow **non-spontaneous**

3) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} > 1$ \rightarrow **spontaneous**

4) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} < 1$ \rightarrow **non-spontaneous**

Ex 2. The ΔH° for the reaction $2\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}_2(l) + \text{H}_2(g)$ is $+136.3 \text{ kJ}$

The following reactions are spontaneous

$\text{H}_2\text{O}_2(l) \rightarrow \text{H}_2\text{O}(l) + \frac{1}{2}\text{O}_2(g)$ $\Delta H^\circ = -98.0 \text{ kJ}$

$\text{H}_2(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$ $\Delta H^\circ = -285.8 \text{ kJ}$

Calculate the ΔH° for the reaction of H_2O_2 to H_2 and O_2

$\text{H}_2\text{O}_2(l) \rightarrow \text{H}_2(g) + \text{O}_2(g)$ $\Delta H^\circ = ?$

$\Delta H^\circ_{rxn} = 136.3 \text{ kJ}$ $\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$ $\Delta H^\circ_{rxn} = -285.8 \text{ kJ}$

1) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} > 1$

Ex 3. Calculate ΔH°_{rxn} for the following reaction at 25°C

$\text{P}_4\text{O}_{10}(s) + 6\text{H}_2\text{O}(l) \rightarrow 4\text{H}_3\text{PO}_4(aq)$

$\Delta H^\circ_f(\text{H}_3\text{PO}_4(aq)) = -1279 \text{ kJ/mol}$ $\Delta H^\circ_f(\text{P}_4\text{O}_{10}(s)) = -2912 \text{ kJ/mol}$

$\Delta H^\circ_{rxn} = -1279 \text{ kJ}$ $\Delta H^\circ_{rxn} = -2912 \text{ kJ}$ **$\Delta H^\circ_{rxn} = -1279 \text{ kJ}$** $\Delta H^\circ_{rxn} = -2912 \text{ kJ}$

Ex 4. Calculate the heat of formation of $\text{H}_2\text{O}(l)$ from the combustion data given below

$\text{H}_2(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$

Heat of formation
 $\text{H}_2(g)$ 0 kJ/mol
 $\text{O}_2(g)$ 0 kJ/mol

$\Delta H^\circ_{rxn} = -285.8 \text{ kJ}$ $\Delta H^\circ_{rxn} = -285.8 \text{ kJ}$ $\Delta H^\circ_{rxn} = -285.8 \text{ kJ}$ $\Delta H^\circ_{rxn} = -285.8 \text{ kJ}$

Ex 5. Calculate the standard heat of formation ΔH°_f of $\text{H}_2\text{O}(l)$ from the following information:

$\text{H}_2\text{O}_2(l) \rightarrow \text{H}_2\text{O}(l) + \frac{1}{2}\text{O}_2(g)$ $\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$

$\Delta H^\circ_f(\text{H}_2\text{O}_2(l)) = -187.8 \text{ kJ/mol}$

$\Delta H^\circ_f(\text{H}_2\text{O}(l)) = -285.8 \text{ kJ/mol}$

$\Delta H^\circ_f(\text{H}_2\text{O}(l)) = -285.8 \text{ kJ/mol}$

$\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$ $\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$ $\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$ $\Delta H^\circ_{rxn} = -98.0 \text{ kJ}$

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