

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solubility _____ is the ability of a substance to dissolve in another substance. A substance that does not dissolve is (5) _____ insoluble _____.

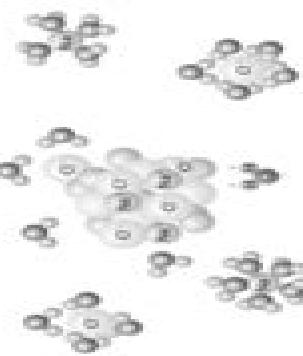
(6) _____ liquid _____ solubility of the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, water is the (7) _____ solute _____ and oil is the (8) _____ insoluble _____.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.



1. Indicate in solute is when the solute is in a solution. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Section 15.2 **solvent****STUDY GUIDE FOR CONTENT MASTERY****Section 15.3** **solvent****solvent**

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a test tube with the test of solution may be dissolved, stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvent, with test of solute.

True _____ 16. The overall energy change that occurs when a solute dissolves in a solvent is called the heat of solution.

False _____ 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False _____ 18. Solubility continues to change as the solute dissolves in a solvent because the heat of solution is lost.

True _____ 19. In a saturated solution, solute and crystallization are in equilibrium.

True _____ 20. Additional solute can be dissolved in an unsaturated solution.

True _____ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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 - The Rise of Digital Reading Study Guide Answers Content Mastery
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Answers Content Mastery
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Answers Content Mastery
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Answers Content Mastery
 - Personalized Recommendations

- Study Guide Answers Content Mastery User Reviews and Ratings
- Study Guide Answers Content Mastery and Bestseller Lists

5. Accessing Study Guide Answers Content Mastery Free and Paid eBooks

- Study Guide Answers Content Mastery Public Domain eBooks
- Study Guide Answers Content Mastery eBook Subscription Services
- Study Guide Answers Content Mastery Budget-Friendly Options

6. Navigating Study Guide Answers Content Mastery eBook Formats

- ePUB, PDF, MOBI, and More
- Study Guide Answers Content Mastery Compatibility with Devices
- Study Guide Answers Content Mastery Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Study Guide Answers Content Mastery
- Highlighting and Note-Taking Study Guide Answers Content Mastery
- Interactive Elements Study Guide Answers Content Mastery

8. Staying Engaged with Study Guide Answers Content Mastery

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Answers Content Mastery

9. Balancing eBooks and Physical Books Study Guide Answers Content Mastery

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Study Guide Answers Content Mastery

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Study Guide Answers Content Mastery

- Setting Reading Goals Study Guide Answers Content Mastery
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide Answers Content Mastery

- Fact-Checking eBook Content of Study Guide Answers Content Mastery

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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