

Chapter 11: The Behavior of Gases – Study Guide

Test Date: _____

- The four variables that affect gases:
 - Pressure
 - be able to convert between pressure units of torr, mm Hg, kPa, and atm
(*Note: the relationships between units will be given to you on the test as follows: 1 atm = 760 mm Hg = 760 torr = 101.325 kPa*)
 - Temperature – ALWAYS in kelvin!!!!
 - you **must memorize** the conversion between Celsius and Kelvin ($K = ^\circ C + 273$)
 - Volume
 - Moles
- Boyle's law – pressure & volume change (indirectly related)
 - Hint: Vice President Boyle
- Charles's law – volume & temperature change (directly related)
 - Hint: Charlie Brown is on I.V.
- Gay-Lussac's law – pressure & volume change (directly related)
 - Hint: In the Lou, you need I.P.
- Combined gas law
- Ideal gas law
 - $PV = nRT$ (be able to solve for any unknown variable)
 - Also be able to solve for moles using $PV = nRT$ & convert moles to mass using the molar mass of the gas (or vice-versa)
 - For this equation you must use one the specific units listed below:
 - Pressure – atm, Volume – liters, Temperature – kelvin, n – moles, R – ideal gas constant (*Note: the value of "R" and its units will be given to you on the test as follows: $0.0821 \text{ L}\cdot\text{atm}/(\text{mol}\cdot\text{K})$*)
- The value of standard temperature and pressure conditions (STP) **must be memorized**
 - Standard temperature = 0°C or 273 K
 - Standard pressure = 1 atm
- Dalton's Law of Partial Pressure
 - In a mixture of gases, the total pressure of the mixture is equal to the sum of the partial pressures of each gas in the mixture ($P_t = P_1 + P_2 + P_3 \dots$ etc)
 - When a gas is **collected over water**, Dalton's Law looks like $P_{\text{total}} = P_{\text{gas}} + P_{\text{H}_2\text{O}}$
 - The partial pressure of the water vapor ($P_{\text{H}_2\text{O}}$) at a specific temperature can be found on the water vapor table (or sometimes just listed in the problem)

Note: The equations for the gas laws will be given on the test but you must be able to identify each type of gas law problem by name (Boyle's, Charles's, Gay-Lussac's, combined, ideal.

Study Guide Behavior Of Gases Answer Key

Didier Musso



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