

"Mastering Chemical Equilibrium: A Comprehensive Study Guide"

Chemical Equilibrium Study Guide:

I. Introduction:

- Definition: State where the rate of the forward and reverse reactions are equal.
- Significance: Understanding chemical equilibrium helps in explaining the behavior of reactions under different conditions.

II. Factors Affecting Chemical Equilibrium:

- Concentration of reactants and products
- Temperature
- Pressure
- Catalyst

III. Equilibrium Constant (Kc):

- Definition: Measure of the relative concentrations of reactants and products at equilibrium.
- Calculation: $K_c = [C]^c [D]^d / [A]^a [B]^b$
- Significance: K_c can be used to predict the direction of the reaction under different conditions.

IV. Le Chatelier's Principle:

- Definition: States that a system at equilibrium will counteract any disturbance to maintain the balance.
- Application: Can be used to predict the effect of changes in concentration, temperature,

Study Guide Chemical Equilibrium

Julie R. Frentrup

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