

## Unit 5: Chemical Quantities and Reactions

### **Chapter 10 – Chemical Quantities, Chapter 11 – Chemical Reactions, Ch. 12 Stoichiometry**

#### **Objectives:**

1. Demonstrate the three methods for measuring the amount of a substance – count, mass, and volume.
2. Understand the concept of a mole and what it represents.
3. Convert between the mass of a substance, the number of moles that mass contains, and the number of particles in that mass.
4. Calculate the volume of a gas at STP.
5. Demonstrate how chemical equations describe chemical reactions.
6. Illustrate how to balance chemical reactions by changing coefficients.
7. State and describe the five general types of reactions.
8. Given the equation of a reaction, classify the reaction as one of the general types.
9. Interpret balanced chemical equations and understand all the information present in them.
10. Perform mole-mole, mass-mass, mass-particles, volume-volume, and other stoichiometry calculations.
11. Determine the amount of product produced in a chemical reaction when taking into consideration a limiting reagent.
12. Determine the percent yield of a chemical reaction.

#### **Vocabulary:**

Mole (noun)  
Mol (unit)  
Avogadro's Number  
Molar Mass (g/mol)  
STP  
Molar Volume (22.4L/mol)  
Reactants  
Products  
Word equation  
Chemical Equation  
Coefficient  
Combination / Synthesis Reaction

Decomposition Reaction  
Single-replacement / displacement Reaction  
Activity Series  
Double-replacement / displacement Reaction  
Combustion Reaction  
Mole Ratio  
Limiting Reagent  
Excess Reagent  
Theoretical Yield  
Actual Yield  
Percent Yield

#### 10.1 The Mole: A Measurement of Matter

Read p. 287-296 and make your own notes by answering the following questions.  
How do we measure amounts?

**Stoichiometry** is a method used to measure the amounts of substances involved in chemical reactions and relate them to one another. It comes from Greek and Old English words meaning a series of steps for measuring something. For example, a sample's mass or volume can be converted to the number of particles in the sample. Since atoms are so small, in order to count how many are in a sample, we group them into larger units. The group or unit of measurement used to count number of atoms, molecules, or formula units of substances is the **mole**, abbreviated **mol**. It is most convenient to define the mole as the number equal to the number of carbon atoms in exactly 12 grams of pure carbon-12. The number of things in one mole is  $6.02 \times 10^{23}$ , which is known as **Avogadro's number**. The mass of one mol of an **element** is called the **molar mass**, where 1amu is equivalent to 1mol. So the mass of 1mol of carbon is 12g since the atomic mass of carbon is 12amu. The units for molar mass are grams per mol, g/mol.

# Study Guide Chemical Quantities

**Michael Rajnik**

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## **Table of Contents Study Guide Chemical Quantities**

1. Understanding the eBook Study Guide Chemical Quantities
  - The Rise of Digital Reading Study Guide Chemical Quantities
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Chemical Quantities
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Chemical Quantities
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Chemical Quantities
  - Personalized Recommendations
  - Study Guide Chemical Quantities User Reviews and Ratings
  - Study Guide Chemical Quantities and Bestseller Lists
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  - Study Guide Chemical Quantities Enhanced eBook Features
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  - Adjustable Fonts and Text Sizes of Study Guide Chemical Quantities
  - Highlighting and Note-Taking Study Guide Chemical Quantities
  - Interactive Elements Study Guide Chemical Quantities
8. Staying Engaged with Study Guide Chemical Quantities
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Chemical Quantities
9. Balancing eBooks and Physical Books Study Guide Chemical Quantities
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Chemical Quantities
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Chemical Quantities
  - Setting Reading Goals Study Guide Chemical Quantities
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Chemical Quantities
  - Fact-Checking eBook Content of Study Guide Chemical Quantities
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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