

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

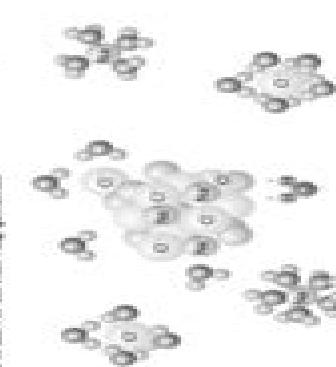
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.

1. In what is solute in which the solute is a solid? What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Section 15.2 **STUDY GUIDE FOR CONTENT MASTERY****Section 15.1** **Answers**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecule around the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In order to separate the ions of solute may be dissolved.

Starting at shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with type of solute.

True 14. The overall ionic charge that occurs when a solute dissolves in a solvent is lost of solutes.

False 15. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False 16. Solvent molecules is large as the solute but is less than the solute.

True 17. In a saturated solution, solute and solvents are in equilibrium.

True 18. Additional solute can be dissolved in an unsaturated solution.

True 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Study Guide For Content Master 15 Key

Albert Kaskel

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Tintinalli's Emergency Medicine: A Comprehensive Study Guide, 9th edition Judith E. Tintinalli,O. John

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Psychology Catalog 2005 Neil Thomson, 2004-09

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Table of Contents Study Guide For Content Master 15 Key

1. Understanding the eBook Study Guide For Content Master 15 Key
 - The Rise of Digital Reading Study Guide For Content Master 15 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Master 15 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Master 15 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Master 15 Key
 - Personalized Recommendations
 - Study Guide For Content Master 15 Key User Reviews and Ratings
 - Study Guide For Content Master 15 Key and Bestseller Lists
5. Accessing Study Guide For Content Master 15 Key Free and Paid eBooks
 - Study Guide For Content Master 15 Key Public Domain eBooks
 - Study Guide For Content Master 15 Key eBook Subscription Services
 - Study Guide For Content Master 15 Key Budget-Friendly Options
6. Navigating Study Guide For Content Master 15 Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - Study Guide For Content Master 15 Key Compatibility with Devices
 - Study Guide For Content Master 15 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Content Master 15 Key
 - Highlighting and Note-Taking Study Guide For Content Master 15 Key
 - Interactive Elements Study Guide For Content Master 15 Key

8. Staying Engaged with Study Guide For Content Master 15 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Master 15 Key
9. Balancing eBooks and Physical Books Study Guide For Content Master 15 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Master 15 Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Master 15 Key
 - Setting Reading Goals Study Guide For Content Master 15 Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide For Content Master 15 Key
 - Fact-Checking eBook Content of Study Guide For Content Master 15 Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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 - Interactive and Gamified eBooks

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