

### Lesson 3: Limiting reactants

- Reactions stop when one of the reactants is used up.

Limiting reactant - is used completely in the reaction; determines the amount of product formed

Excess reactant - leftover once the reaction stops.

- To calculate which reactant is limiting, you will compare the amount of each reactant to a single product (the true amount of product formed came from the limiting)



If 200.0 g of sulfur reacts with 100.0 g of chlorine, what mass of disulfur dichloride is produced? Which is the limiting reactant?

given = 200.0 g S<sub>8</sub>; 100.0 g Cl<sub>2</sub>  
unknown = ? g S<sub>2</sub>Cl<sub>2</sub>; ? limiting reactant

$$200.0 \text{ g S}_8 \times \frac{1 \text{ mol S}_8}{256.528 \text{ g S}_8} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{1 \text{ mol S}_8} \times \frac{135.038 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \\ = 421 \text{ g S}_2\text{Cl}_2$$

$$100.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{4 \text{ mol S}_2\text{Cl}_2}{4 \text{ mol Cl}_2} \times \frac{135.038 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} = \\ = \boxed{190 \text{ g S}_2\text{Cl}_2}$$

Chlorine is the limiting reactant



If 100.0 g Na and 100.0 g Fe<sub>2</sub>O<sub>3</sub> react, what mass of Iron is produced? Which is the limiting reactant?

given = 100.0 g Na; 100.0 g Fe<sub>2</sub>O<sub>3</sub>  
unknown = ? g Fe; ? limiting reactant

$$100.0 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.990 \text{ g Na}} \times \frac{2 \text{ mol Fe}}{6 \text{ mol Na}} \times \frac{55.845 \text{ g Fe}}{1 \text{ mol Fe}} = \\ = 81.0 \text{ g Fe}$$

$$100.0 \text{ g Fe}_2\text{O}_3 \times \frac{1 \text{ mol Fe}_2\text{O}_3}{159.687 \text{ g Fe}_2\text{O}_3} \times \frac{2 \text{ mol Fe}}{1 \text{ mol Fe}_2\text{O}_3} \times \frac{55.845 \text{ g Fe}}{1 \text{ mol Fe}} = \\ = \boxed{70.0 \text{ g Fe}}$$

L.R. = Fe<sub>2</sub>O<sub>3</sub>

# Study Guide For Content Mastery Limiting Reactants

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In a world driven by information and connectivity, the power of words has become more evident than ever. They have the ability to inspire, provoke, and ignite change. Such could be the essence of the book **Study Guide For Content Mastery Limiting Reactants**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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