

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide For First Year College Chemistry

David Kirk



Study Guide For First Year College Chemistry:

Modular Study Guide for First Semester General Chemistry (CHM 1045) Marta E. Goicoechea-Pappas, Anthony J. Pappas, Miami Dade College. Kendall Campus. Chemistry Department, 2006 **Resources in Education** ,1984 *The United States Catalog* Mary Burnham, Carol Hurd, 1928 **DOD Pam** United States. Office of Armed Forces Information and Education, **United States Armed Forces Institute Catalog** United States Armed Forces Institute, 1963 The United States Catalog ,1924 **Scientific and Technical Books and Serials in Print** ,1984 *El-Hi Textbooks and Serials in Print* , 1995 R R Bowker Publishing, 1995-02 Chemist and Druggist ,1890 **General College Chemistry** Colorado Agricultural and Mechanical College. Department of Chemistry, Paul Reheard Frey, 1952 *County Agent's Magazine* ,1923 Announcements Florida State University. Summer Session, 1917 **El-Hi Textbooks in Print** ,1984 The Publishers Weekly ,1926 **British Medical Journal** ,1894 Study Guide for Life on Earth Carl S. Pike, 1973 **Catalog and Announcement of North Dakota Agricultural College for ...** ,1907 **Quarterly Guide for Readers** Finsbury (England). Public Library, 1914 **Science** ,1892 Vols for 1911 13 contain the Proceedings of the Helminothological Society of Washington ISSN 0018 0120 1st 15th meeting **Pamphlets on Forestry in Washington** ,1908

The book delves into Study Guide For First Year College Chemistry. Study Guide For First Year College Chemistry is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Study Guide For First Year College Chemistry, encompassing both the fundamentals and more intricate discussions.

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- Chapter 2: Essential Elements of Study Guide For First Year College Chemistry
- Chapter 3: Study Guide For First Year College Chemistry in Everyday Life
- Chapter 4: Study Guide For First Year College Chemistry in Specific Contexts
- Chapter 5: Conclusion

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4. In chapter 3, this book will examine the practical applications of Study Guide For First Year College Chemistry in daily life. The third chapter will showcase real-world examples of how Study Guide For First Year College Chemistry can be effectively utilized in everyday scenarios.
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6. In chapter 5, the author will draw a conclusion about Study Guide For First Year College Chemistry. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Study Guide For First Year College Chemistry.

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