

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

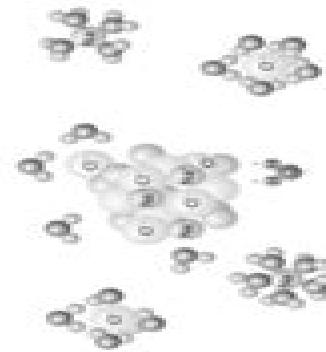
solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

A(n) **10. solution** of sugar is dissolved in water.The oxygen atom in the **11. water** molecule is the **12. oxygen** atom.

13. solute because it goes into solution in a solvent. In **14. water** is an **15. soluble** substance. A substance that does not dissolve is **16. insoluble**. In **17. liquid** solvents in the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, water is a **18. soluble** liquid and oil is a **19. insoluble** liquid.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.

1. In water, a solute is in which the solute is in water? What is a solute?

Solutions is the process of surrounding solute particles with solvent particles to form a solution.

Study Guide for Content Mastery

Oxidation States and Change • Chapter 11

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Section 15.1 content**STUDY GUIDE FOR CONSTITUTANT MASTERY**

16. In sodium chloride dissolved in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are **separated and surrounded by the water molecules.**

17. Explain the attraction of the water molecule around the sodium ion and chloride ion.
Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?
The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a test tube with the test of solution, stirring the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with true or false.

True **16.** The overall energy change that occurs when a solute dissolves is called the heat of solution.

False **17.** Solubility is a property of the solute, amount of water the solute is in a given amount of solution is a property of temperature and pressure.

False **18.** Solvent molecules in a liquid solution are in constant motion. **19.** Hydration is a test of solubility.

True **20.** In a saturated solution, solute and crystallization are in equilibrium.

True **21.** Additional solute can be dissolved in an unsaturated solution. **22.** The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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