

Equilibrium Study Guide



- Consider the equilibrium above. Which of the following changes will increase the concentration of HgI_4^{2-} ?
 - (A) Increasing the concentration of OH^{-}
 - (B) Adding 6 M HNO_3
 - (C) Increasing the mass of HgO present
 - (D) Increasing the temperature
 - (E) Adding a catalyst
- In which of the following systems would the number of moles of substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?
 - (A) $\text{CO(g)} + \text{NO(g)} \leftrightarrow \text{CO}_2(\text{g}) + \frac{1}{2} \text{N}_2(\text{g})$
 - (B) $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \leftrightarrow 2 \text{NH}_3(\text{g})$
 - (C) $\text{N}_2(\text{g}) + 2 \text{O}_2(\text{g}) \leftrightarrow 2 \text{NO}_2(\text{g})$
 - (D) $\text{N}_2\text{O}_4(\text{g}) \leftrightarrow 2 \text{NO}_2(\text{g})$
 - (E) $\text{NO(g)} + \text{O}_3(\text{g}) \leftrightarrow \text{NO}_2(\text{g}) + \text{O}_2(\text{g})$
- Which of the following is true for a redox reaction that is thermodynamically favorable at standard conditions?
 - (A) $\Delta G^{\circ} > 0$, $K_{\text{eq}} > 1$, $E^{\circ}_{\text{cell}} > 0$
 - (B) $\Delta G^{\circ} > 0$, $K_{\text{eq}} > 1$, $E^{\circ}_{\text{cell}} < 0$
 - (C) $\Delta G^{\circ} < 0$, $K_{\text{eq}} > 1$, $E^{\circ}_{\text{cell}} > 0$
 - (D) $\Delta G^{\circ} < 0$, $K_{\text{eq}} < 1$, $E^{\circ}_{\text{cell}} < 0$
 - (E) $\Delta G^{\circ} < 0$, $K_{\text{eq}} < 1$, $E^{\circ}_{\text{cell}} > 0$
- For the reaction $2 \text{A} \leftrightarrow 4 \text{B}$, $K_c = 2.0$. Determine K_c for $4 \text{B} \leftrightarrow 2 \text{A}$.
 - (A) 2.0
 - (B) 4.0
 - (C) 0.50
 - (D) 0.20
 - (E) 0.40

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