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 $\text{CaH}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$
 $\text{CO}_2 + 10\text{H}_2\text{O} + 4\text{KBr} \rightarrow 2\text{K}_2\text{CO}_3 + 8\text{H}_2\text{O} + 4\text{KBr}$
 $\text{Na}_2\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2$
 SnF_2

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $\text{C}_6\text{H}_{10}\text{O}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_4$
 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{Na}_2\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2$
 SnF_2
 $\text{N}_2\text{O} + \text{CO}_2 \rightarrow \text{NaBr}$
 solve the following stoichiometry grams grams problems 6

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $\text{C}_6\text{H}_{10}\text{O}_2 + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $4\text{KBr} + 3\text{Na}_2\text{SO}_3 \rightarrow 2\text{K}_2\text{CO}_3 + 8\text{H}_2\text{O} + 4\text{KBr}$
 $5\text{Na}_2\text{SO}_3 + 3\text{H}_2\text{O} \rightarrow 2\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$

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 $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
 note all reactants and products require a coefficient of at least one stuck review related articles

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