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Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$q = ?$$

$$\Delta t = 55 - 22$$

$$= 33^{\circ}\text{C}$$

$$m = 10\text{ g}$$

$$S = 0.9 \text{ J/g°C}$$

$$q = m \cdot S \cdot \Delta t$$

$$= 10 \times 0.9 \times 33$$

$$= 297 \text{ J}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\Delta t = 26.4 - 24.5 = 2$$

$$S = 4.18 \text{ J/g°C}$$

$$m = 335 \text{ g}$$

$$q = m \cdot S \cdot \Delta t$$

$$= 335 \times 4.18 \times 2$$

$$= 2800.6 \text{ J}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\Delta t = 25 - 10 = 15$$

$$m = 225 \text{ g}$$

$$S = 4.18 \text{ J/g°C}$$

$$q = m \cdot S \cdot \Delta t$$

$$= 225 \times 4.18 \times 15$$

$$= 14107.5 \text{ J}$$

$$1 \text{ J} = 1 \times 10^{-3} \text{ KJ}$$

$$\therefore 14.1075 \text{ KJ}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$m = 1 \text{ kg} = 1000 \text{ g}$$

$$q = 1000 \times 4.18 \times 74$$

$$\Delta t = 99 - 25 = 74$$

$$S = 4.18 \text{ J/g°C}$$

$$q = 309320 \text{ J}$$

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal? $S = ?$

$$C = m \cdot S \rightarrow S = \frac{C}{m}$$

$$(q_{\text{water}} = q_{\text{metal}})$$

$$q_{\text{water}} = 75 \times 4.18 \times (28.34 - 24)$$

$$= 1252$$

$$q = m \cdot S \cdot \Delta t$$

$$1252 = 26 \times S \times (82.25 - 28.34)$$

$$1252 = 26 \times 63.91 \times S$$

$$S = 0.893 \text{ J/g°C}$$

4. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$C = 1265 \text{ J/g°C}$$

$$\Delta t = 25.12 - 22.34 = 2.78$$

$$q = ?$$

$$q = C \cdot \Delta t$$

$$= 1265 \times 2.78$$

$$= 3516.1 \text{ J}$$

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