

Chemistry 101 Final Exam Study Guide- Spring 2017

Acids and Bases

Acid: substance that provides H^+ ions in water

Base: substance that provides OH^- ions in water

Covalent Bonds

- Review: Ionic bonds are formed by the positive charge of the cation being attracted to the negative charge of the anion
- Covalent bonds: bonds formed by sharing electrons between the atoms (not ions)
 - 1 electron from each atom (2 atoms per bond)
 - Electrons available for bonding are the un-paired electrons in orbitals (orbitals with only 1 electron in them)
- Molecule
 - A group of atoms held together by covalent bonds
- Atoms will bond with enough other atoms to complete the octet
 - But they can only form as many bonds as unpaired valence electrons
 - Hydrogen has 1 valence electron, so it can only form 1 bond, and will not achieve an octet
 - Boron has 3 valence electrons, so can only form 3 bonds, and will not achieve the octet (only 6 electrons)
 - There are always exceptions to the rules
 - Sulfur and phosphorus can extend the octet

Multiple Covalent Bonds between 2 atoms

- If the octet is not met, additional bonds may be formed
 - CO_2 (carbon dioxide)
 - C has 4 valence electrons, so 4 bonds possible
 - O has 6 valence electrons, 2 are unpaired, so 2 bonds are possible

PRACTICE

IONS:

- Sodium
 - Symbol: Na
 - Valence: 1
 - Electron dot: Na^+
 - Ionic dot: Na^+ (loses its electron to get a $1+$ charge)
- Magnesium
 - Symbol: Mg
 - 2 valence E-
 - Loses 2 electrons to get $2+$ charge
- Oxygen
 - Symbol: O
 - Valence: 6
 - Will gain 2 electrons for a $2-$ charge

Cannot have a covalent bond with ions, just atoms

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Technical Preceptors into the Classroom Mary Poulton and John Kemeny 17 Academic Excellence Workshops Boosting Success in Technical Courses Ruth A Streveler 18 Supplemental Instruction at an Urban Community College Joyce Ship Zaritsky 19 Peer Assisted Teaching and Learning in Distance Education Judith A Couchman 20 Using Structured Study Groups To Create Chemistry Honors Sections Brian P Coppola Douglas S Daniels and Jason K Pontrello 21 Student Mentoring and Community in a University Honors Program Ronald E Mickel 22 Where Undergraduates Are the Experts Peer Based Instruction in the Writing Center Dennie Paoli and Eric Hobson 23 Peer Facilitators of In Class Groups Adapting Problem Based Learning to the Undergraduate Setting Deborah E Allen and Harold B White iii 24 Student Directed Instruction in an Undergraduate Psychopathology Course Cheryl Golden and Calverta McMorris 25 Peer Writing Tutors Lisa Lebduska 26 The Workshop Project Peer Led Team Learning in Chemistry Jerry L Sarquis Linda J Dixon David K Gosser Jack A Kampmeier Vicki Roth Victor S Strosak and Pratibha Varma Nelson 27 a Introductory Psychology Laboratory Designed and Taught by Undergraduate Teaching Interns Stephen P Stelzner Michael G Livingston and Thomas Creed 28 Undergraduate Teaching Assistants Bring Active Learning to Class Melissa A Thibodeau 29 Student Faculty Partnerships To Develop Teaching and Enhance Learning Milton D Cox 30 Educating the Critic Student Driven Quality Elizabeth Kinland Lisa Firing Lenze Lynn Melendez Moore and Larry D Spence and 31 College Teachers and Student Consultants Collaborating about Teaching and Learning D Lynn Sorenson Four appendixes contain examples of hiring documents training syllabi teaching materials and evaluation procedural documents Contains 18 figures 59 tables and 178 references Sld

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Table of Contents Spring Final Exam Study Guide Chemistry

1. Understanding the eBook Spring Final Exam Study Guide Chemistry
 - The Rise of Digital Reading Spring Final Exam Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Spring Final Exam Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Spring Final Exam Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Spring Final Exam Study Guide Chemistry
 - Personalized Recommendations
 - Spring Final Exam Study Guide Chemistry User Reviews and Ratings
 - Spring Final Exam Study Guide Chemistry and Bestseller Lists
5. Accessing Spring Final Exam Study Guide Chemistry Free and Paid eBooks

- Spring Final Exam Study Guide Chemistry Public Domain eBooks
- Spring Final Exam Study Guide Chemistry eBook Subscription Services
- Spring Final Exam Study Guide Chemistry Budget-Friendly Options

6. Navigating Spring Final Exam Study Guide Chemistry eBook Formats

- ePUB, PDF, MOBI, and More
- Spring Final Exam Study Guide Chemistry Compatibility with Devices
- Spring Final Exam Study Guide Chemistry Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Spring Final Exam Study Guide Chemistry
- Highlighting and Note-Taking Spring Final Exam Study Guide Chemistry
- Interactive Elements Spring Final Exam Study Guide Chemistry

8. Staying Engaged with Spring Final Exam Study Guide Chemistry

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Spring Final Exam Study Guide Chemistry

9. Balancing eBooks and Physical Books Spring Final Exam Study Guide Chemistry

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Spring Final Exam Study Guide Chemistry

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Spring Final Exam Study Guide Chemistry

- Setting Reading Goals Spring Final Exam Study Guide Chemistry
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Spring Final Exam Study Guide Chemistry

- Fact-Checking eBook Content of Spring Final Exam Study Guide Chemistry
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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