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George T. Geis



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web if a solute is soluble in a solvent the overall process is said to be spontaneous if a solute does not dissolve in a solvent insoluble the solution process is said to be non spontaneous a spontaneous process or change proceeds in a given direction without needing to be driven by an external energy input

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web solubility and solubility product mh 5 16 1 16 2 in this section we are going to consider the solubility of ionic solids in water solubility may be considered to be an equilibrium the equilibrium is between solid and ions in solution any ionic solid is 100% ionized in aqueous solution once it actually dissolves

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web information solubility rules a solute is considered soluble if an appreciable amount of it can be dissolved in a given amount of the solvent for example both table salt NaCl and table sugar $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ are soluble substances in water a solute is considered insoluble if very little of it dissolves in a given amount of the solvent

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web predict what might happen to the solubility of the insoluble salt if a 1.0 M acid solution was added drop wise to the beaker hint which ion species in the beaker will react with the acid how will this reaction alter the concentration of the common ion