

- b) If the rate of consumption of magnesium is 5.0×10^{-9} mol/s, find the rate of consumption of HCl in moles/s.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{2 \text{ mol HCl}}{1 \text{ mol Mg}} = 1.0 \times 10^{-8} \text{ mol HCl/s}$$

Answer $1.0 \times 10^{-8} \text{ mol HCl/s}$

- c) If the rate of consumption of magnesium is 5.0×10^{-9} mol/s, find the rate of production of H₂ in g/s.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} =$$

Answer $1.0 \times 10^{-8} \text{ g H}_2/\text{s}$

- d) If the rate of consumption of magnesium is 5.0×10^{-9} mol/s, find the rate of production of H₂ in L/s (@STP).

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} =$$

Answer $1.1 \times 10^{-7} \text{ L H}_2/\text{s}$

- c) If the rate of consumption of magnesium is 5.0×10^{-9} mol/s, find the mass of Mg consumed in 5.0 minutes.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{24.3 \text{ g Mg}}{1 \text{ mol Mg}} \times \frac{60}{1} \times 5.0 \text{ min} = 3.6 \times 10^{-5} \text{ g Mg}$$

Answer $3.6 \times 10^{-5} \text{ g Mg}$

3. When butane (C₄H₁₀) is burned in air (oxygen), the products carbon dioxide and water are formed.

- a) Write a balanced formula equation for this reaction.



- b) If butane is consumed at an average rate of 0.116 grams/s, determine the rate of production of CO₂ in g/s.

$$0.116 \frac{\text{g C}_4\text{H}_{10}}{\text{s}} \times \frac{1 \text{ mol C}_4\text{H}_{10}}{58.0 \text{ g C}_4\text{H}_{10}} \times \frac{8 \text{ mol CO}_2}{2 \text{ mol C}_4\text{H}_{10}} \times \frac{44.0 \text{ g CO}_2}{1 \text{ mol CO}_2} = 0.352 \text{ g CO}_2/\text{s}$$

Answer $0.352 \text{ g CO}_2/\text{s}$

Stoichiometric Calculations Answer Key

George Tyler Miller

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