

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffmann apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before	2 mol	0 mol	0 mol
Change	-2 mol	+2 mol	
After	0 mol	2 mol	

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before	9.94 mol	xs mol	0 mol
Change	-9.94 mol		+9.94 mol
After	0 mol	9.94 mol	

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before	1.0 mol	xs mol	0 mol	0 mol
Change	-1.0 mol		-1.0 mol	
After	0 mol	xs mol		

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_{6} \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_{6}}{180 \text{ g C}_6\text{H}_{12}\text{O}_{6}} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_{6} \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_{6}} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

# Stoichiometry Guide Answer Key

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Smith, Jack Price**

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**General Chemistry Key Concise Review Guide Sixth Edition** Darrell D. Ebbing, 1999     **The Complete Idiot's Guide to Chemistry** Ian Guch, 2006 With topics such as stoichiometry chemical equilibria and thermodynamics it's no wonder so many students have a bad reaction to chemistry class Fortunately The Complete Idiot's Guide to Chemistry Second Edition gives readers an easy to understand formula for success when it comes to passing their chemistry class with flying colours Intended for both high school students who are required to take a college prep chemistry course and college students who need to take chemistry to meet their degree requirements this book gives readers a non-intimidating easy to understand companion to their textbooks The book follows a standard math based chemistry curriculum covering the following topics

Atomic structure and elements The periodic table Ionic and covalent compounds Chemical reactions Kinetics Acids and bases  
Electrochemistry Organic chemistry Nuclear chemistry Stoichiometry Thermodynamics *Stoichiometry* B. I. Bhatt,1996

Resources in Education ,1980 **Study Guide to Accompany Basics for Chemistry** Martha Mackin,1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Chemistry, Study Guide* James E. Brady,John R. Holum,1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples *Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney* William H. Myers,William T. Mooney,1980 *Holt Chemistry* Ralph Thomas Myers,2004 *Student Solutions Manual: Ssm Chemistry* Deborah Wiegand,2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems *Michigan Test for Teacher Certification Study Guide ,2001* Prentice Hall Chemistry ,2000 **Study Guide and Problems Book for Biochemistry, Garrett and Grisham** David Karl Jemiolo, Garrett, Charles M. Grisham,1996 **Merrill Chemistry** Robert C. Smoot,Smoot, Richard G. Smith,Jack Price,1998 *Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl* Paul B. Kelter,1993 *General Chemistry* B. Richard

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