

Stoichiometry Study Guide KEY

Chemistry

RHS – Mr. Moss

1. Define the following:
 - a. *Stoichiometry—the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.*
 - b. *Mole – The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.*
 - c. *Mole Ratio – The ratio between any two substances in a balanced chemical equation.*
 - d. *Molar Mass – The mass in grams of one mole of any substance.*
2. What relationships can be determined from a balanced chemical equation? *The relationships between particles, moles, and mass for all reactants and products.*
3. Explain why mole ratios are central to stoichiometric calculations. *Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.*
4. What is the first step in all stoichiometric calculation? *(Be careful.) Write a balanced chemical equation.*
5. How is molar mass used in some stoichiometric calculations? *Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.*
6. Find the molar mass of the following:
 - a. H_2SO_4 (sulfuric acid) 98.09 g/mol
 - b. ZnO (zinc oxide) 81.39 g/mol
 - c. NH_4NO_3 (ammonium nitrate) 80.05 g/mol
7. List the mole ratios of the following *balanced* chemical equation.
 - a. $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$ $\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} : \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3} : \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$ *(Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.)*

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John A. Olmsted, Gregory M. Williams

Stoichiometry Study Guide Key:

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