

Balancing Single Replacement Reactions

In a single replacement reaction, the cation (usually a metal) in a compound is replaced by another, more reactive metal. They always follow this format:



As an example, look at this chemical equation for single replacement reaction between solid lithium and a solution of iron (II) nitrate.



Lithium replaced iron (II) in the compound. Iron was left alone as a solid product. This reaction, however, is not balanced. Count the atoms. You might find it helpful to treat the nitrate polyatomic ion as a single element in balancing.

Reactants

$$\text{Li} = \underline{\hspace{2cm}}$$

$$\text{Fe} = \underline{\hspace{2cm}}$$

$$\text{NO}_3 = \underline{\hspace{2cm}}$$

Products

$$\text{Li} = \underline{\hspace{2cm}}$$

$$\text{Fe} = \underline{\hspace{2cm}}$$

$$\text{NO}_3 = \underline{\hspace{2cm}}$$

Add coefficients to balance the equation and satisfy the Law of Conservation of Matter.



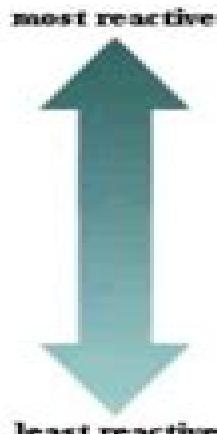
Not all combinations of elements and compounds in single replacement reactions will actually react! You can use the activity series of metals to predict this. Follow these two rules:

1. *Highly reactive metals prefer to be part of a compound.*
2. *Highly stable metals prefer to be alone.*

Look at the following combinations. If they will react, then predict what the products will be. If they will not react, then write NR for "no reaction".



lithium (Li)
potassium (K)
barium (Ba)
calcium (Ca)
sodium (Na)
magnesium (Mg)
aluminum (Al)
zinc (Zn)
cobalt (Co)
nickel (Ni)
lead (Pb)
copper (Cu)
silver (Ag)
gold (Au)



Single Replacement Reactions Study Guide

ML Morrison

Single Replacement Reactions Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

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Study Guide/Selected Solutions

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