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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal KBr that

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the O_2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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