



# Solubility Rules

Rule #		SOLUBILITY RULES
1	Cations	ALL Group 1 alkali metal ( $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , $\text{Rb}^+$ , $\text{Cs}^+$ , $\text{Fr}^+$ ) and Ammonium ( $\text{NH}_4^+$ ) compounds are <i>soluble</i> in water. Most, but NOT all, Group 2 alkali earth metal compounds are <i>soluble</i> .
2	Halides	Most Halides are <i>soluble</i> . Chlorides ( $\text{Cl}^-$ ), Bromides ( $\text{Br}^-$ ), and Iodides ( $\text{I}^-$ ) are <i>soluble</i> EXCEPT Silver ( $\text{Ag}^+$ ), Mercury ( $\text{Hg}_2^{2+}$ ), and Lead ( $\text{Pb}^{2+}$ )
3	Soluble Anions	Nitrates ( $\text{NO}_3^-$ ), Acetates ( $\text{C}_2\text{H}_3\text{O}_2^-$ ), Chlorate ( $\text{ClO}_3^-$ ), and Perchlorates ( $\text{ClO}_4^-$ ) are <i>soluble</i>
4	Sulfates	Most Sulfates ( $\text{SO}_4^{2-}$ ) are <i>soluble</i> EXCEPT compounds of Barium ( $\text{Ba}^{2+}$ ), Strontium ( $\text{Sr}^{2+}$ ), Calcium ( $\text{Ca}^{2+}$ ), Lead ( $\text{Pb}^{2+}$ )
5	Insoluble Anions	Most Phosphates ( $\text{PO}_4^{3-}$ ), Carbonates ( $\text{CO}_3^{2-}$ ), Sulfides ( $\text{S}^{2-}$ ), & Sulfites ( $\text{SO}_3^{2-}$ ) are <i>insoluble</i> EXCEPT for compounds of Group 1 alkali metals and Ammonium ( $\text{NH}_4^+$ )
6	Hydroxides	Most hydroxides ( $\text{OH}^-$ ) are <i>insoluble</i> EXCEPT compounds of Group 1 alkali metals, most Group 2 alkali earth metals, and Ammonium ( $\text{NH}_4^+$ )

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