

NAME: _____

Molarity Practice

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1. Calculate the molarity of the following solutions:

a) 45.0 g of NaOH in 2.00 L of solution

b) 10.0 g of CaCl_2 in 250 mL of solutionc) 25.0 g of CuSO_4 in 500 mL of solution

2. What is the molarity of a solution that contains 10.0 g of each of the following?

a) 10.0 g of NaOH

b) 10.0 g of NaCl

3. What is the molarity of a solution that contains 10.0 g of each of the following in 250 mL of solution?

4. Which of the following contains the highest concentration? Show using calculations.

a) 10.0 g of NaOH in 250 mL of solution

b) 20.0 g of NaOH in 500 mL

5. Make a more concentrated solution of sodium hydroxide (NaOH) by adding to a solution that has a concentration of 1.0 M. What you need to make 1.0 L of a 2.0 M solution? (Use molarity, the molarity and the formula, and the molar mass of sodium hydroxide.)

Skill Practice 53 Answers Chemistry

Darrell D. Ebbing



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