

## Solubility Rules Worksheet

1. Classify each substance as being soluble or insoluble in water.

- a. Potassium bromide – \_\_\_\_\_ f. Silver iodide - \_\_\_\_\_  
b. Lead (II) carbonate – \_\_\_\_\_ g. Cadmium (II) sulfide - \_\_\_\_\_  
c. Barium sulfate – \_\_\_\_\_ h. Zinc carbonate - \_\_\_\_\_  
d. Zinc hydroxide – \_\_\_\_\_ i. Silver acetate - \_\_\_\_\_  
e. Sodium acetate – \_\_\_\_\_ j. Copper (II) sulfide – \_\_\_\_\_

2. Identify the two new compounds which form if the solutions were mixed. Circle any compounds that would precipitate from the solution (are insoluble).

	KBr	$\text{Na}_2\text{CO}_3$	CaS	$\text{NH}_4\text{OH}$
$\text{AgNO}_3$				
$\text{BaCl}_2$				
$\text{Al}(\text{NO}_3)_3$				
$\text{CuSO}_4$				

3. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formu	Name	Solubility
$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$		
	Barium hydroxide	
$\text{FeCO}_3$		
NaOH		
	Rubidium nitrate	
$\text{MgSO}_4$		
	Zinc chloride	
	Zinc hydroxide	

# Solubility Rules Lab Answer Key

**Thomas Griffiths**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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