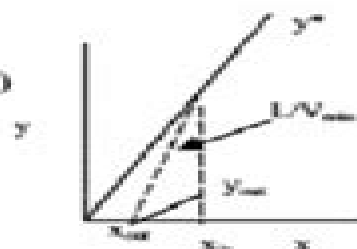


18.12

(a)



$$M_{\text{toluene}} = 92.13$$

$$y^* = P_{\text{tol}}/P = 256x / 1.1 = 233 x$$

1 liter has $1000 / 18 = 55.56 \text{ g mol H}_2\text{O}$
at 60 ppm, toluene concentration

$$= \frac{60 \times 10^{-6} \times 1000}{92.13} = 6.51 \times 10^{-4} \text{ g mol / liter}$$

$$x_m = \frac{6.51 \times 10^{-4}}{55.56} = 1.17 \times 10^{-5}$$

$$\Delta x = 58 / 60 \times x_m = 1.13 \times 10^{-5}$$

By material balance, $L \Delta x = V \Delta y$

at V_{min} , $\Delta y = y^*$ for x_m

$$V_{\text{min}} = \frac{L \Delta x}{233 x_m} = \frac{L}{233} \times \frac{58 \text{ ppm}}{60 \text{ ppm}} = 4.15 \times 10^{-3} L$$

$$L = 100 \text{ l / min} \times 55.56 \text{ mol / l} = 5556 \text{ mol / min}$$

$$V_{\text{min}} = 5556 \times 4.15 \times 10^{-3} = 23.06 \text{ mol / min}$$

(b) at $V = 2 V_{\text{min}} = 46.11 \text{ mol / min}$

$$y_m = \frac{L \Delta x}{V} = \frac{5556 \times 1.13 \times 10^{-5}}{46.11} = 1.36 \times 10^{-5}$$

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