

### 1.1\*. Concentration of a gas mixture.

A mixture of noble gases [helium (1), argon (2), krypton (3), and xenon (4)] is at a total pressure of 200 kPa and a temperature of 400 K. If the mixture has equal mole fractions of each of the gases, determine:

a) The composition of the mixture in terms of mass fractions.

Solution:

Basis: 100 kmole of the mixture

ORIGIN := 1

$$\text{moles} := \begin{pmatrix} 25 \\ 25 \\ 25 \\ 25 \end{pmatrix} \text{ kmole}$$

kmole := 1000 · mole

$$\text{MW} := \begin{pmatrix} 4 \\ 40 \\ 83.8 \\ 131.3 \end{pmatrix} \frac{\text{kg}}{\text{kmole}}$$

mass :=  $\overrightarrow{(\text{moles} \cdot \text{MW})}$

$$\text{mass} := \begin{pmatrix} 100 \\ 1 \times 10^{-3} \\ 2.095 \times 10^{-3} \\ 3.283 \times 10^{-3} \end{pmatrix} \text{ kg}$$

$$\text{total} := \sum_{i=1}^4 \text{mass}_i$$

$$\text{total} = 6.478 \times 10^3 \text{ kg}$$

$$\text{mass\_fraction} := \frac{\overrightarrow{\text{mass}}}{\text{total}}$$

$$\text{mass\_fraction} := \begin{pmatrix} 0.015 \\ 0.154 \\ 0.323 \\ 0.507 \end{pmatrix}$$

b) The average molecular weight of the mixture.

Solution:

kmole := 1000 · mole

$$M_{av} := \frac{\text{total}}{100 \cdot \text{kmole}}$$

$$M_{av} = 0.065 \frac{\text{kg}}{\text{mol}}$$

c) The total molar concentration.

Solution:

T := 400 K

1kPa := 1000 Pa

$$P := 200 \text{ kPa}$$

$$R := 8.314 \frac{\text{Pa} \cdot \text{m}^3}{\text{kmole} \cdot \text{K}} = 1 \frac{\text{kg}}{\text{mol} \cdot \text{K}}$$

$$c := \frac{P}{R \cdot T} \quad c = 60.14 \frac{\text{mol}}{\text{m}^3}$$

d) The mass density.

Solution:

$$\rho := c \cdot M_{av}$$

$$\rho = 3.896 \frac{\text{kg}}{\text{m}^3}$$

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