

Unit 3 Review - Atomic Structure & Periodic Table

Part 1: Answer the following questions.

1. a. Which scientist created the first modern atomic theory? John Dalton
 b. What was his theory?
 1. All matter is made of extremely small particles called atoms
 2. All atoms of a given element are identical and atoms of a specific element are different from any other element
 3. Atoms cannot be created, divided into smaller particles or destroyed
 4. Different atoms combine in simple whole number ratios to form compounds
 5. In a chemical reaction, atoms are separated, combined or rearranged.
- c. What was his model of the atom called? Billiard Ball
- d. What was proven wrong and why?
 1. All atoms of a given element are identical – wrong b/c of the existence of isotopes
 2. Atoms cannot be divided into smaller particles – wrong b/c of the existence of electrons, protons and neutrons
2. a. Which scientist did the gold foil experiment? Rutherford
 b. What did he conclude from the gold foil experiment?
 1. Atoms are made up of mostly empty space
 2. Atoms have a small positive center call the nucleus
3. a. Who discovered the electron? J. J. Thomson
 b. What was the name of his experiment? Cathode Ray Tube Experiment
 c. What was his model of the atom called? Plum Pudding Model
4. What did Mendeleev do?
Arranged the periodic table in order of increasing atomic mass
5. a. Who arranged the **modern** periodic table? Moseley
 b. How is the modern periodic table arranged? Arranged in order of Atomic Number
6. a. What are the three subatomic particles? Give the symbols, charge, mass and location of each.

Particles	Symbol	Charge	Mass	Location
Proton	p^+	Positive	1 amu	Nucleus
Neutron	n^0	Neutral	1 amu	Nucleus
Electron	e^-	Negative	1/2000 amu	Electron Cloud

7. What is special about the proton? **The proton IDENTIFIES THE ELEMENT**
8. Fill in the blanks in the following:
 - a. Mass # = $p^+ + n^0$
 - b. Atomic # = # of protons
9. What is the difference between Average Atomic mass and mass #? (Average) Atomic Mass is found on the periodic table and is the weighted average of the isotopes. The Mass Number is the mass of one element and must be given or calculated by adding the protons and neutron of that element.
10. Which of the following are isotopes of the same element? How do you know? A, and D. They have the same Atomic Number (# of Protons)
 - a. $^{12}_{5}\text{X}$
 - b. $^{15}_{5}\text{X}$
 - c. $^{12}_{6}\text{X}$
 - d. $^{15}_{5}\text{X}$
11. Calculate the average atomic mass for Boron if it has two naturally occurring isotopes: boron -10 (abundance of 19.8%) and boron - 11 (abundance of 80.2%).

$$\{ 10 \times \frac{19.8}{100} \} + \{ 11 \times \frac{80.2}{100} \} = 10.802 \text{ amu}$$
12. Identify an element "X" by calculating the average atomic mass. X = 6 (abundance of 7.5%) and X= 7 (abundance 92.5%).

$$\{ 6 \times \frac{7.5}{100} \} + \{ 7 \times \frac{92.5}{100} \} = 6.925 \text{ amu} \quad \text{Which is Lithium}$$

Study Guide Nuclear Chemistry Answer Key

**Denise Guinn, Rebecca Brewer, Rachel
C. Lum**



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1. Understanding the eBook Study Guide Nuclear Chemistry Answer Key
 - The Rise of Digital Reading Study Guide Nuclear Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Nuclear Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Nuclear Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Nuclear Chemistry Answer Key
 - Personalized Recommendations

- Study Guide Nuclear Chemistry Answer Key User Reviews and Ratings
- Study Guide Nuclear Chemistry Answer Key and Bestseller Lists
- 5. Accessing Study Guide Nuclear Chemistry Answer Key Free and Paid eBooks
 - Study Guide Nuclear Chemistry Answer Key Public Domain eBooks
 - Study Guide Nuclear Chemistry Answer Key eBook Subscription Services
 - Study Guide Nuclear Chemistry Answer Key Budget-Friendly Options
- 6. Navigating Study Guide Nuclear Chemistry Answer Key eBook Formats
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 - Study Guide Nuclear Chemistry Answer Key Compatibility with Devices
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 - Adjustable Fonts and Text Sizes of Study Guide Nuclear Chemistry Answer Key
 - Highlighting and Note-Taking Study Guide Nuclear Chemistry Answer Key
 - Interactive Elements Study Guide Nuclear Chemistry Answer Key
- 8. Staying Engaged with Study Guide Nuclear Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Nuclear Chemistry Answer Key
- 9. Balancing eBooks and Physical Books Study Guide Nuclear Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Nuclear Chemistry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Nuclear Chemistry Answer Key
 - Setting Reading Goals Study Guide Nuclear Chemistry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Nuclear Chemistry Answer Key
 - Fact-Checking eBook Content of Study Guide Nuclear Chemistry Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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