

Study Buddy: Classification of Chemical Reactions

Terms	Definitions	Representation									
Law of Conservation of Mass	During a chemical or physical change, mass is conserved. Mass is neither created nor destroyed. Mass of reactants equals mass of products.	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ <table border="1"> <tr> <td>N_2</td><td>3H_2</td><td>$\rightarrow 2\text{NH}_3$</td></tr> <tr> <td>28g</td><td>6g</td><td>34g</td></tr> <tr> <td>2 atoms</td><td>6 atoms</td><td>8 atoms</td></tr> </table>	N_2	3H_2	$\rightarrow 2\text{NH}_3$	28g	6g	34g	2 atoms	6 atoms	8 atoms
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Precipitation Reaction	Occurs when two aqueous solutions react to produce a solid precipitate.	$\text{NaCl}_{(aq)} + \text{AgNO}_{3(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{3(aq)}$									
Redox Reaction	Occurs when one or more electrons is transferred during a chemical reaction. Reduction: occurs when a species gains an electron Oxidation: occurs when a species loses an electron	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$ $\text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^-$									

Representations:

- Balance the following chemical equations and draw a particle diagram showing conservation of mass:

- Hydrogen gas and oxygen gas react to form liquid water.



- Solid zinc is added to an aqueous solution of silver nitrate to produce zinc nitrate and solid silver.



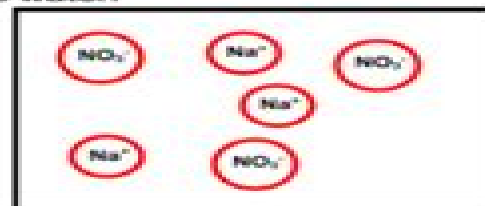
- A chemist added a 5.0g piece of solid magnesium metal to 10.0g of hydrochloric acid ($\text{HCl}_{(aq)}$) in a 50 mL beaker. During the reaction, the solution in the beaker became warmer and produced many bubbles. Once the reaction finished, the mass of the final solution was only 14.6g. What is the best explanation for the "missing" mass?

A bubbles are an indicator that a gas is being produced. In this case, 0.4g of gas were produced and no longer in the solution.

0.21 mole of Mg added to excess HCl produces 0.21 moles of $\text{H}_{2(g)}$ or 0.4g of $\text{H}_{2(g)}$

- Describe how you would make 100. mL of 2.0M NaNO_3 from powdered NaNO_3 . In the box, show what happens to the NaNO_3 when it is added to water.

- Determine mass of NaNO_3 needed:
 $2.0\text{M} = (\text{X g}/85.0 \text{ g/mol})/0.100 \text{ L}$ $\text{X} = 17\text{g}$
- Put 17g of NaNO_3 in a 100 mL volumetric flask
- Add water to 100 mL
- Mix



*Water molecules not shown for clarity of ions

Study Guide Redox Reaction

E Staff



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Table of Contents Study Guide Redox Reaction

1. Understanding the eBook Study Guide Redox Reaction
 - The Rise of Digital Reading Study Guide Redox Reaction
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Redox Reaction
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Redox Reaction
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Redox Reaction
 - Personalized Recommendations
 - Study Guide Redox Reaction User Reviews and Ratings
 - Study Guide Redox Reaction and Bestseller Lists

5. Accessing Study Guide Redox Reaction Free and Paid eBooks
 - Study Guide Redox Reaction Public Domain eBooks
 - Study Guide Redox Reaction eBook Subscription Services
 - Study Guide Redox Reaction Budget-Friendly Options
6. Navigating Study Guide Redox Reaction eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Redox Reaction Compatibility with Devices
 - Study Guide Redox Reaction Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Redox Reaction
 - Highlighting and Note-Taking Study Guide Redox Reaction
 - Interactive Elements Study Guide Redox Reaction
8. Staying Engaged with Study Guide Redox Reaction
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Redox Reaction
9. Balancing eBooks and Physical Books Study Guide Redox Reaction
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Redox Reaction
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Redox Reaction
 - Setting Reading Goals Study Guide Redox Reaction
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Redox Reaction
 - Fact-Checking eBook Content of Study Guide Redox Reaction
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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