

# Gas Laws Study Guide

Lesson 1: Gas Laws	
<p>1. <b>Boyle's Law:</b> If the volume of a gas is decreased, the pressure increases. If the volume of a gas is increased, the pressure decreases.</p> <p>2. <b>Charles' Law:</b> If the temperature of a gas is increased, the volume increases. If the temperature of a gas is decreased, the volume decreases.</p> <p>3. <b>Avogadro's Law:</b> If the number of molecules of a gas is increased, the volume increases. If the number of molecules of a gas is decreased, the volume decreases.</p> <p>4. <b>Combined Gas Law:</b> <math>P_1V_1/T_1 = P_2V_2/T_2</math></p> <p>5. <b>Gay-Lussac's Law:</b> <math>P_1/T_1 = P_2/T_2</math></p> <p>6. <b>Avogadro's Hypothesis:</b> <math>PV = nRT</math></p>	<p><b>Gas Laws Study Guide</b></p> <p>Boyle's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Charles' Law: <math>V_1/T_1 = V_2/T_2</math></p> <p>Avogadro's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Combined Gas Law: <math>P_1V_1/T_1 = P_2V_2/T_2</math></p> <p>Gay-Lussac's Law: <math>P_1/T_1 = P_2/T_2</math></p> <p>Avogadro's Hypothesis: <math>PV = nRT</math></p> <p>Gas Law Equations:</p> <p>Boyle's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Charles' Law: <math>V_1/T_1 = V_2/T_2</math></p> <p>Avogadro's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Combined Gas Law: <math>P_1V_1/T_1 = P_2V_2/T_2</math></p> <p>Gay-Lussac's Law: <math>P_1/T_1 = P_2/T_2</math></p> <p>Avogadro's Hypothesis: <math>PV = nRT</math></p> <p>Gas Law Equations:</p> <p>Boyle's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Charles' Law: <math>V_1/T_1 = V_2/T_2</math></p> <p>Avogadro's Law: <math>P_1V_1 = P_2V_2</math></p> <p>Combined Gas Law: <math>P_1V_1/T_1 = P_2V_2/T_2</math></p> <p>Gay-Lussac's Law: <math>P_1/T_1 = P_2/T_2</math></p> <p>Avogadro's Hypothesis: <math>PV = nRT</math></p>
<p>Compare the following statements. Is there a difference? Justify the answer.</p> <p>1. Pressure increases because the molecules are moving faster. <input type="checkbox"/></p> <p>2. Pressure increases because the molecules are moving slower. <input type="checkbox"/></p> <p>3. Pressure increases because the molecules are moving faster. <input type="checkbox"/></p> <p>4. Pressure increases because the molecules are moving slower. <input type="checkbox"/></p> <p>5. The number of molecules of gas in a container increases the volume increases. <input type="checkbox"/></p> <p>6. Pressure increases because the molecules are moving faster. <input type="checkbox"/></p> <p>7. Pressure increases because the molecules are moving slower. <input type="checkbox"/></p>	

# Study Guide The Gas Laws

**Steven S. Zumdahl, Paul B. Kelter**

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