

Single Replacement Reaction

A single replacement reaction occurs when one element replaces another element in a compound.



Examples:



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A single replacement reaction or single displacement reaction is when one element replaces another in a compound.

Get the definition of a single replacement reaction or single displacement reaction. Get examples of single replacement reactions and learn how to use the [metal reactivity series](#) to predict whether a reaction will occur and the products.

Single Replacement Reaction Definition

A **single replacement reaction** is a chemical reaction where one element replaces another in a compound. It is also known as a **single displacement reaction**. The general form of a single replacement reaction chemical equation is:



Single replacement reactions occur when A is more reactive than B or **product** AC is more stable than BC. A and B can be either two metals (including hydrogen; C is [an anion](#)) or else two halogens (C is a cation). If BC and AC are in aqueous solutions, C acts as a spectator ion.

Single Replacement Reaction Examples

There are two different scenarios for single replacement reactions. In one form of the reaction, one cation replaces the other. In the other form of the reaction, one anion replaces the other.

Cation Replacement Examples

Usually the cation is a metal, but it doesn't have to be. Here are examples of single replacement reactions involving the cations:

Single Replacement Reactions Study Guide

William H. Myers, William T. Mooney

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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