

Name: _____

Atomic Structure & the Periodic Table Study Guide

Atomic Structure

Particle	Charge	Where are they found?
Proton	Positive	Nucleus
Neutron	Neutral/ no charge	Nucleus
Electron	negative	Electron cloud/energy levels outside nucleus

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**

2. What are the vertical columns on the periodic table called? **Groups or families**

a. How many of these are on the periodic table? **18**

b. Why are the elements in these columns? **They have the same chemical properties (& same # of valence electrons)**

c. Elements in the same group have the same number of **valence** electrons.

3. What are the horizontal rows on the periodic table called? **periods**

a. How many of these are on the periodic table? **7**

b. Elements in these rows have the same number of **energy (electron shells)** levels.

4. What are valence electrons? **Electrons in the outer energy level**

5. How many electrons are in each energy level?

1st energy level **2** 2nd energy level **8** 3rd energy level **8**

6. The atomic number is equal to the **# of protons and also # of electrons**

7. The atomic mass is equal to the **# protons + # of neutrons**

8. How do you find the number of neutrons? **Atomic mass – atomic number**

9. We view the order of the periodic table today according to the **atomic number**

Study Guide The Periodic Table

Speedy Publishing

Study Guide The Periodic Table:

Periodic Table (Advanced) Speedy Publishing,2014-08-01 An advanced periodic table of elements displays not only the elements but the ions that form each element A pamphlet with such a visual aid would greatly benefit chemistry students Any student taking chemistry will need to learn the elements A pamphlet would be concise and break the information down simply making it easier to understand and remember It allows students to simply focus on the main point rather than taking in information that they may or may not need

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