

Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
2. Stoichiometry is based on the law of conservation of mass.
3. In any chemical reaction, the mass of the products is less than the mass of the reactants.
4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
5. The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.



| Substance | Molar Mass (g/mol) | Number of Molecules | Number of Moles (mol) | Mass (g) |
|-------------------|--------------------|---------------------|-----------------------|----------|
| 6. Methanol | 32.05 | | | |
| 7. Oxygen gas | 32.00 | | | |
| 8. Carbon dioxide | 44.01 | | | |
| 9. Water | 18.02 | | | |

10. What are the reactants? _____
11. What are the products? _____
12. What is the total mass of the reactants? _____
13. What is the total mass of the products? _____
14. How do the total masses of the reactants and products compare? _____

Study Guide Stoichiometry Answer

Leo J. Malone, Theodore O. Dolter

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde

Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s

Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter,Steven S.

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Wingrove,Robert L. Caret,1981 **Study Guide to Accompany Basics for Chemistry** Martha Mackin,1982-01-01 Study

Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each

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problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter

energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on

writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *General Chemistry, Study Guide* James E. Brady,1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems *Study Guide for Chemical Principles* [by] Steven S. Zumdahl Steven S. Zumdahl,Paul B. Kelter,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples **Study Guide for Chemistry by Steven S. Zumdahl** Martha B. Barrett,1986 *Introduction to Chemistry, Study Guide* T. R. Dickson,1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills Many chapters in this edition has been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images **Study Guide/Selected Solutions Manual** Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book **Study Guide** Joan I. Senyk,Darrell D. Ebbing,1984 *Study Guide [to] Chemistry, a Systematic Approach* [by] Sisler, Dresdner, and Mooney William H. Myers,William T. Mooney,1980 **Study Guide to Accompany Chemical Principles, Third Edition** Daniel L. Reger,1985 *Chemistry, Study Guide* James E. Brady,John R. Holom,1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples **Study Guide** Darrell D. Ebbing,Susan M. Schelble,1998 *Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry* Iris Stovall,Donald J. DeCoste,2000 *Study Guide for Chemistry, Third Edition* [by] Steven S. Zumdahl Paul B. Kelter,1993

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