

### 12.1-12.3

What is the **reaction rate**?

the change in concentration of a reactant or product per unit time.

What is an **instantaneous rate**? How is it calculated?

the rate of a chemical reaction at a specific moment in time, using the slope of the tangent line

What is a **rate law**? Be sure to explain **rate constant** and **order**.

an expression that shows how the rate of reaction depends on the concentration of reactants, the rate constant is the proportionality constant in the relationship between reaction rate and reactant concentrations. Order is the exponent, determined by experiment, of the reactant concentration in a rate law.

How does an **integrated rate law** differ from a regular rate law (differential rate law)?

an expression that shows the concentration of a reactant as a function of time while rate law is a reaction as a function of concentration.

What is an **overall reaction order**?

sum of the exponents of the concentration

### 12.4

	First Order	Second Order	Zero Order
Description	directly proportional to the concentration of one reactant.	depends on the square of the concentration of one reactant or the product of the concentrations of two reactants.	independent of the concentration of the reactant(s).
Exponent in a rate law	Rate = $k[A]$	Rate = $k[A]^2$	Rate = $k$
Integrated rate law	$\ln[A] = -kt + \ln[A]_0$	$1/[A] = -kt + 1/[A]_0$	$[A] = -kt + [A]_0$
What makes a straight line?	slope = $-k$	slope = $k$	slope = $-k$
Half life equation	$t_{1/2} = 0.693/k$	$t_{1/2} = 1/k[A]_0$	$t_{1/2} = [A]_0/2k$

# Zumdahl Ch 12 Study Guide

**Steven S. Zumdahl**



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