



## The Mole Study Guide

### Form description

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Avogadro's number refers to the number of particles in one gram of a substance.

- ☐ True
- ☐ False

What is the mass of one mole of aluminum?

- ☐ 26.982 g
- ☐ 13 g
- ☐ 53.965 g
- ☐ 14 g

Which is the correct formula for the compound whose percent composition is shown?

# Study Guide Mole Test

**James E. Brady**



## Study Guide Mole Test:

**Study Guide with Selected Solutions** Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Study Guide/Selected Solutions Manual Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

*Student Study Guide to Accompany Chemistry* Kenneth J. Hughes, Paul B. Kelter, 2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison **Study Guide to**

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