

## Solubility and Concentration

1. Match the following. Write the correct letter in the blank provided.

1. Solute <u>B</u>	<del>A. Tells you how much solvent is dissolved in the solution</del>
2. Solvent <u>E</u>	<del>B. The part that dissolves into the solution</del>
3. Dissolve <u>D</u>	<del>C. Increase solubility by moving molecules faster</del>
4. Stirring <u>C</u>	<del>D. When something is seen to disappear inside the solvent</del>
5. Concentration <u>A</u>	<del>E. The part that dissolves something</del>

2. Determine if the following statements are true or false.

- a. Solubility increases if you stir a solute into a solution. F
- b. All solutes have the same solubility in a given solvent. F
- c. There is a limit on the amount of solute that can dissolve in a given solvent. T
- d. You can dissolve additional solutes in an unsaturated solution. T

3. Give three ways to make sugar dissolve faster. Heat Stirring and increasing surface area

4. Which one will dissolve faster? Circle the correct answer.

- a. Powdered sugar or granulated sugar?
- b. In hot water or in cold water?
- c. Stirred or not stirred?
- d. Large particles or small particles ?

5. 65 g of sugar is dissolved in 750 ml of water. What is the concentration of the solution?  
0.08 g/ml

6. Which is more concentrated: 34 g of salt dissolved in 100 ml of water or 100 g of salt in 1500 ml of water?  
34 g of Salt in 100ml of water

7. If the solubility of salt in water was determined to be 0.5 g/ml, would a solution with 50 g of salt in 150 ml of water be considered saturated?

no

# Solubility And Concentration Answer Key

**Marcia L. Gillette, Ira Blei, Wendy  
Gloffke, George Odian**



## **Solubility And Concentration Answer Key:**

**General, Organic, and Biochemistry Study Guide** Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry

**Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Let's Review Albert S. Tarendash, 1993-09-01 Covers phases of matter atomic structure the chemical bond the periodic table solutions chemical reactions equilibrium acids and bases organic chemistry and lab procedures

*21st Century Chemistry* Kimberley Waldron, 2019-01-17 Waldron 21st Century Chemistry promotes scientific literacy and helps students understand chemistry applications in everyday life With an exceptionally clear and fresh writing style Waldron engages non science majors and provides a focus on environmental topics with Naturebox and Green Beat features Recurring Themes help students remember fundamental take away ideas and concepts so they can apply their knowledge of chemistry as they make choices as consumers voters and overall informed citizens The new second edition of 21st Century Chemistry will include new content featuring fresh stories for roughly four of the Naturebox features and roughly three of the GreenBeats features refreshed end of chapter content including questions encouraging students to research their local environment using web resources media tools focused on a few key resources that address engagement and reading support including videos of current events and real world applications and LearningCurve reading quizzes VitalSource e Book

**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

**Introductory Chemistry** Charles H. Corwin, 2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs

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precise definitions followed by a detailed discussion on prescriptions including their definition essential parts correct handling and common errors Posology is covered with emphasis on factors affecting drug dosage and pediatric dose calculations based on age body weight and body surface area A strong foundation in pharmaceutical calculations is provided through topics like imperial and metric systems of weights and measures percentage solutions alligation proof spirit isotonic solution calculations using freezing point data and molecular weight based problems The text then moves to powders explaining their definition classification advantages disadvantages and the preparation of simple and compound powders dusting powders effervescent efflorescent and hygroscopic powders eutectic mixtures and the principle of geometric dilution Liquid dosage forms are introduced with their advantages disadvantages commonly used excipients and solubility enhancement techniques Monophasic liquids such as gargles mouthwashes throat paints ear drops nasal drops enemas syrups elixirs liniments and lotions are described with their definitions and methods of preparation The book further explains suspensions and emulsions covering their definitions classifications advantages disadvantages methods of preparation types such as flocculated and deflocculated suspensions tests for emulsion type and the various stability problems along with methods to overcome them Suppositories are presented in detail including their definition types advantages disadvantages bases methods of preparation displacement value calculations and evaluation parameters Pharmaceutical incompatibilities are discussed with their definition classification and clear examples of physical chemical and therapeutic incompatibilities that can arise in prescriptions and formulations Finally semisolid dosage forms are covered including their definitions classification mechanisms and factors influencing dermal penetration of drugs preparation of ointments pastes creams and gels excipients used in semisolid formulations and their evaluation Overall the book offers a clear systematic and exam oriented introduction to key concepts of pharmaceutics making it highly suitable for beginner pharmacy students

*Proceedings : Western Regional Meeting* Society of Petroleum Engineers (U.S.) Western Regional Meeting,1997

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Chemical Interactions ,2005

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