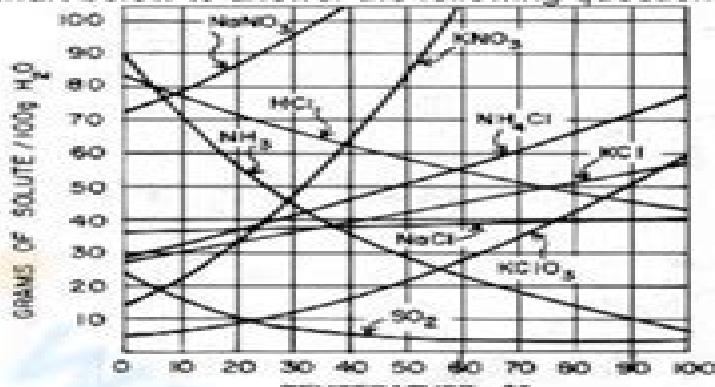


Solubility Curve Worksheet Answers

Use the solubility chart below to answer the following questions:



- What is the solubility of potassium nitrate at 30° C? 46 g / 100 g H₂O
- How many grams of ammonia can I dissolve in 200 grams of water at a temperature of 45° C? 64 grams (2 x 32 grams)
- At what temperature is the solubility of sodium chloride the same as the solubility of potassium chloride? 34° C
- How many grams of ammonium chloride would I need to make 300 grams of a saturated solution at 70° C? 183 grams
- What do all of the compounds that decreased in solubility over the temperature range in the graph have in common? They are all gases
- What compound is least soluble at 40° C? sulfur dioxide
- What ionic compound is least soluble at 40° C? KClO₃

Solubility Model 1 Three Solutions Answers

Mark Steyvers

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Chemistry, an Experimental Science Chemical Education Material Study,1963 *Chemist and Druggist* ,1879
Science I Essential Interactions ,2000-10 *Pharmaceutical Journal* ,1925 **Process Photogram and Illustrator**
,1900 *Chemical Interactions* ,2005 **Transactions of the Pharmaceutical Meetings** ,1926 **The**
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Compounds: Inorganic complexes Consultants Bureau Enterprises,1960* **Chemical Thermodynamics of Solid**
Solutions of Interest in Radioactive Waste Management Organisation for Economic Co-operation and
Development,2007 This volume provides a state of the art report on the modelling of aqueous solid solution systems by the
combined use of chemical thermodynamics and experimental and computational techniques These systems are ubiquitous in
nature and therefore intrinsic to the understanding and quantification of radionuclide containment and retardation processes
present in geological repositories of radioactive waste The various approaches are illustrated with case studies from the
literature **Nature** Sir Norman Lockyer,1908 **Thermodynamic Properties of Nonelectrolyte Solutions** William
Eugene Acree,1984 Thermodynamic Properties of Nonelectrolyte Solutions reviews several of the more classical theories on
the thermodynamics of nonelectrolyte solutions Basic thermodynamic principles are discussed along with predictive methods
and molecular thermodynamics This book is comprised of 12 chapters the first of which introduces the reader to
mathematical relationships such as concentration variables homogeneous functions Euler s theorem exact differentials and
method of least squares The discussion then turns to partial molar quantities ideal and nonideal solutions and empirical
expression **Journal of Gas Lighting and Water Supply** ,1907 *Chemistry in the Community (ChemCom)* American
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