

SOLUBILITY STUDY GUIDE- Multiple Choice Section

Multiple Choice Section: This study guide is a compilation of questions from provincial exams since 2000. I urge you to become intimately familiar with question types. You will notice that questions from one year to another are very similar in their composition. Identification of question types will allow you to be more efficient in answering these questions on the provincial examination. My recommendations for using this study guide are as follows:

1. **DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet.
2. **RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
3. **LEARN FROM YOUR MISTAKES!** If you get a question wrong, **figure out why**. If you are having difficulty, **TALK TO YOUR STUDY PARTNER**, or maybe **ask your teacher** or **ask your peers**. **Group discussion**. Get together with group members, or other students from class, and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it). Ask your teacher to explain the questions to you during tutorial or after school. **Never ever should be to use** **ANSWER KEYS** or **ANSWER BOOKLETS** to determine **RIGHT** - learning from your mistakes in this booklet will really help you in your efforts to meet this goal.
4. **This is REALLY CRITICAL**. **DO NOT** mark the answer **on the questions themselves**. For example, do not circle any of options A, B, C or D inside the question box. Instead, use a different sheet of paper to place your answers on. By avoiding this urge, you can re-use this study guide effectively again, when preparing for your final exam. In the box to the left, put an asterisk or small note to yourself to indicate that you got the question wrong and need to come back to it. If you got the question correct initially, a check mark might be assistance that you understand this type of question and the notes can concentrate on other questions that present a challenge to you.
5. **Check off the STATUS box on the PRESCRIBED LEARNING OUTCOMES**. I have tried to organize the questions in the identical sequence to which they appear on your Acid Base Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the **UNIT EXAM** and **PROVINCIAL EXAM**!

TABLE OF CONTENTS:

CONCEPT OF SOLUBILITY	1
SOLUBILITY AND PRECIPITATION	5
QUANTITATIVE ASPECTS	15
ANSWERS	21

CONCEPT OF SOLUBILITY

1. **Q1** Which of the following will dissolve in water to produce a molecular solution?
A. CaCl2 B. NaOH C. CH3OH D. Sn(OH)3
2. **Q1** Which one of the following would form an ionic solution when dissolved in water?
A. I2 B. CH3OH C. Ca(NO3)2 D. C12H22O11
3. **Q1** Molecular solutions do not conduct electricity because they contain:
A. molecules only. B. cations and anions. C. molecules and anions. D. molecules and cations.
4. **Q1** When dissolved in water, which of the following produces an ionic solution?
A. O2 B. CH4 C. CaCl2 D. C12H22O11
5. **Q1** When dissolved in water, which of the following forms a molecular solution?
A. HCl(aq) B. NaNO3(aq) C. CH3OH(aq) D. K2SO4(aq)
6. **Q1** Which of the following dissolves in water to form an ionic solution?
A. O2 B. SiO2 C. KMnO4 D. C12H22O11
7. **Q1** Which of the following produces a molecular solution when dissolved in water?
A. CaCl2 B. CH3OH C. NH4SCN D. NaCH3COO
8. **Q1** Consider the following solutes:

I.	<chem>K3PO4</chem>
II.	<chem>C2H5OH</chem>
III.	<chem>C12H22O11</chem>
IV.	<chem>KCH3COO</chem>

Which of the solutes above form only molecular aqueous solutions?
A. I and II
B. II and III
C. II, III and IV
D. I, II, III and IV

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Sabine Zange

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"O" Level Study Guide - Chemistry Quite Easily Done , Fundamentals of Medicinal Chemistry and Drug Metabolism

M. O. Faruk Khan, Ashok Philip, 2018-06-01 The primary objective of this 4 volume book series is to educate PharmD students on the subject of medicinal chemistry. The book set serves as a reference guide to pharmacists on aspects of chemical basis of drug action. This first volume of the series is comprised of 8 chapters focusing on basic background information about medicinal chemistry. It takes a succinct and conceptual approach to introducing important fundamental concepts required for a clear understanding of various facets of pharmacotherapeutic agents, drug metabolism and important biosynthetic pathways that are relevant to drug action. Notable topics covered in this first volume include the scope and importance of medicinal chemistry in pharmacy education, a comprehensive discussion of the organic functional groups present in drugs and information about four major types of biomolecules: proteins, carbohydrates, lipids, nucleic acids and key heterocyclic ring systems. The concepts of acid-base chemistry and salt formation and their applications to the drug action and design follow thereafter. These include concepts of solubility and lipid water partition coefficient (LWPC), isosterism, stereochemical properties, mechanisms of drug action, drug receptor interactions, critical for pharmacological responses of drugs and much more. Students and teachers will be able to integrate the knowledge presented in the book and apply medicinal chemistry concepts to understand the pharmacodynamics and pharmacokinetics of therapeutic agents in the body.

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Table of Contents Solubility Study Guide Multiple Choice Section

1. Understanding the eBook Solubility Study Guide Multiple Choice Section
 - The Rise of Digital Reading Solubility Study Guide Multiple Choice Section
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Study Guide Multiple Choice Section
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Study Guide Multiple Choice Section
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Study Guide Multiple Choice Section
 - Personalized Recommendations
 - Solubility Study Guide Multiple Choice Section User Reviews and Ratings
 - Solubility Study Guide Multiple Choice Section and Bestseller Lists
5. Accessing Solubility Study Guide Multiple Choice Section Free and Paid eBooks
 - Solubility Study Guide Multiple Choice Section Public Domain eBooks
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 - Solubility Study Guide Multiple Choice Section Budget-Friendly Options
6. Navigating Solubility Study Guide Multiple Choice Section eBook Formats

- eBook, PDF, MOBI, and More
- Solubility Study Guide Multiple Choice Section Compatibility with Devices
- Solubility Study Guide Multiple Choice Section Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Solubility Study Guide Multiple Choice Section
- Highlighting and Note-Taking Solubility Study Guide Multiple Choice Section
- Interactive Elements Solubility Study Guide Multiple Choice Section

8. Staying Engaged with Solubility Study Guide Multiple Choice Section

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Solubility Study Guide Multiple Choice Section

9. Balancing eBooks and Physical Books Solubility Study Guide Multiple Choice Section

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Solubility Study Guide Multiple Choice Section

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Solubility Study Guide Multiple Choice Section

- Setting Reading Goals Solubility Study Guide Multiple Choice Section
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Solubility Study Guide Multiple Choice Section

- Fact-Checking eBook Content of Solubility Study Guide Multiple Choice Section
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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