

CHAPTER 1 | Matter and Energy: An Atomic Perspective

1.1. Collect and Organize

Figure P1.1(a) shows “molecules,” each consisting of one red sphere and one blue sphere, and Figure P1.1(b) shows separate blue spheres and red spheres. For each image we are to identify the class of matter. We need to determine whether the substance(s) depicted is/are solid, liquid, or gas and whether the images show an element, a compound, or a mixture that is homogeneous or heterogeneous.

Analyze

An element is composed of all the same type of atoms, and a compound is composed of two or more types of atoms. Solids have a definite volume and a highly ordered arrangement where the particles are close together. Liquids also have a definite volume but have a disordered arrangement of particles that are close together. Gases have disordered particles that fill the volume of the container and are far apart from each other. Homogeneous mixtures have a uniform distribution and composition, and heterogeneous mixtures contain regions of different composition.

Solve

- Because each particle in Figure P1.1(a) consists of one red sphere and one blue sphere, all the particles are the same—that is a compound. The particles fill the container and are disordered, so those particles are in the gas phase.
- Because it shows a mixture of red and blue spheres, Figure P1.1(b) depicts a mixture of blue elemental atoms and red elemental atoms. The blue spheres fill the container and are disordered, so those particles are in the gas phase. The red spheres have a definite volume and are ordered, so those particles are in the solid phase. This is a heterogeneous mixture.

Think About It

Remember that both elements and compounds may be either pure or present in a mixture.

1.2. Collect and Organize

Figure P1.2(a) shows “atoms” of only red spheres, and Figure P1.2(b) has “molecules” consisting of two red spheres or two blue spheres. For each image we are to identify the class of matter. We need to determine whether the substance(s) depicted is/are solid, liquid, or gas and whether the images show an element, a compound, or a mixture that is homogeneous or heterogeneous.

Analyze

A pure substance (whether element or compound) is composed of all the same type of molecule or atom, not a mixture of two kinds. An element is composed of all the same type of atom, and a compound is composed of two or more types of atoms. Solids have a definite volume and a highly ordered arrangement where the particles are close together. Liquids also have a definite volume but have a disordered arrangement of particles that are close together. Gases have disordered particles that fill the volume of the container and are far apart from each other. Homogeneous mixtures have a uniform distribution and composition, and heterogeneous mixtures contain regions of different composition.

Solve

- Because all the atoms are of the same type, Figure P1.2(a) depicts a pure element. The particles take up a definite volume and are ordered, so that element is in the solid phase.
- Because it shows a mixture of blue diatomic molecules and red diatomic molecules, Figure P1.2(b) depicts a mixture of two elements. Both the blue and red diatomic particles fill the container’s volume and are highly disordered; the mixture depicted is in the gas phase. This is a homogeneous mixture.

Think About It

Elements do not need to be present as single atoms. They may be diatomic, as in H₂ or Br₂, or even more highly associated, as in S₈ or P₄.

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