

TYPES OF CHEMICAL REACTIONS LAB PART#1

I. Purpose: To view the actual chemical reactions, write the correct balanced chemical equation, and type of chemical reaction.

Connection to Business: Businesses utilize many types of chemical reactions to produce, maintain, and sterilize their product or to prevent pollution.

II. Materials: burner, test tube, metal tongs (only...no rubber ends), microspatula, evaporating dish, wood splints

Chemicals: copper (wire), magnesium (ribbon), copper (II) carbonate
SAFETY GOGGLES!

III. Procedure: Part #1

1. Obtain materials.
2. Use the fine sandpaper to clean a piece of copper wire until shiny.
3. Set up a Bunsen burner.
4. Using tongs, hold the wire in the hottest part of the burner flame for 1-2 minutes. Record your observation of the wire's appearance afterwards.
5. Place an evaporating dish near the Bunsen burner and repeat step #4 except place the remaining magnesium into the evaporating dish.
(WARNING: AVOID LOOKING DIRECTLY AT THE FLAME)
6. Place 2 heaping microspatulas of copper (II) carbonate into a clean, dry test tube.
7. Using a test tube holder, hold the test tube (at a 45° angle pointed away from people) over the Bunsen burner flame for 3 minutes or until completely black.
(Allow this to cool slightly before placing into the wood test tube holder.)
8. Light a wood splint and place inside the test tube. Record your observations.
9. Dispose of all the metals and burnt powders in a zip-lock bag and place in the trash.
10. Clean and return all lab equipment.

IV. Data/Questions: Part #1 Data Table

Sample	Appearance Before Reaction	Appearance After Reaction
1. Copper	Answer: shiny, reddish brown	Answer: black, powder coating
2. Magnesium	Answer: silver, tinsel	Answer: black ashes, bright light
3. Mercury (II) Oxide <i>(will not perform experiment)</i>	Reddish powder	Liquid mercury formed; flame burned brighter when placed in the test tube
4. Copper (II) Carbonate	Answer: green powder	Answer: decomposes into gas CO ₂ and black residue

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Table of Contents Types Of Reactions Lab

1. Understanding the eBook Types Of Reactions Lab
 - The Rise of Digital Reading Types Of Reactions Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Types Of Reactions Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Types Of Reactions Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Types Of Reactions Lab
 - Personalized Recommendations
 - Types Of Reactions Lab User Reviews and Ratings
 - Types Of Reactions Lab and Bestseller Lists

5. Accessing Types Of Reactions Lab Free and Paid eBooks
 - Types Of Reactions Lab Public Domain eBooks
 - Types Of Reactions Lab eBook Subscription Services
 - Types Of Reactions Lab Budget-Friendly Options
6. Navigating Types Of Reactions Lab eBook Formats
 - ePUB, PDF, MOBI, and More
 - Types Of Reactions Lab Compatibility with Devices
 - Types Of Reactions Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Types Of Reactions Lab
 - Highlighting and Note-Taking Types Of Reactions Lab
 - Interactive Elements Types Of Reactions Lab
8. Staying Engaged with Types Of Reactions Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Types Of Reactions Lab
9. Balancing eBooks and Physical Books Types Of Reactions Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Types Of Reactions Lab
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Types Of Reactions Lab
 - Setting Reading Goals Types Of Reactions Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Types Of Reactions Lab
 - Fact-Checking eBook Content of Types Of Reactions Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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