

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of  $K_a$  for the acid HBrO. (4 marks)  $pH = 4.65$  so  $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



[I]	0.25	0	0
[C]	$-2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]} = \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$ 

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $pH = 2.355$   $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



[I]	$C_0$	0	0
[C]	$-0.004416$	$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$	$0.004416$	$0.004416$

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)} = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 \leftarrow \\ = 0.30439 M$$

Answer  $C_0 = 0.30 M$ 

5. Find the value of  $K_b$  for the oxalate ion ( $C_2O_4^{2-}$ ). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

Answer  $K_b = 1.6 \times 10^{-10}$ 

6. The value of  $K_b$  for the weak base methylamine ( $CH_3NH_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of  $K_a$  for the acid  $CH_3NH_3^+$ . (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

Answer  $K_a = 2.3 \times 10^{-11}$

# Stoichiometry Study Guide Answer Key Pearson

**Paula Yurkanis Bruice**

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