

The number of moles of substances, when compared and reduced to the lowest terms, yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:


2:1:2

1:1:1

4:3:2

1:2:1:1

2:1:2

2:1:1

Convert the following masses to moles by using the information on your periodic table.

a) 34.0 g  $\text{NH}_3$  FW = 17.03 g/mol

$$34.0\text{g} \left( \frac{1\text{mol}}{17.0\text{g}} \right) = 2.00\text{mol}$$

2.00 mol

b) 51.4 g  $\text{Na}_2\text{O}$  FW = 62.0 g/mol

$$51.4\text{g} \left( \frac{1\text{mol}}{62.0\text{g}} \right) = 0.829\text{mol}$$

0.829 mol

c) 4.50 g  $\text{H}_2\text{SO}_4$  FW = 98.1 g/mol

$$4.50\text{g} \left( \frac{1\text{mol}}{98.1\text{g}} \right) = 0.0459\text{mol}$$

0.0459 mol

d) 229 g  $\text{Ca}(\text{NO}_3)_2$  FW = 164.1 g/mol

$$229\text{g} \left( \frac{1\text{mol}}{164.1\text{g}} \right) = 1.40\text{mol}$$

1.40 mol

e) 79.5 g  $\text{NH}_4\text{OH}$  FW = 35.0 g/mol

$$79.5\text{g} \left( \frac{1\text{mol}}{35.0\text{g}} \right) = 2.27\text{mol}$$

2.27 mol

Convert the following to grams.

a) 17.1 moles  $\text{H}_2\text{O}$  FW = 18.0 g/mol

$$17.1\text{mol} \times \left( \frac{18.0\text{g}}{1\text{mol}} \right) = 308\text{g}$$

308 g.

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**Camilla Rothe**



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