

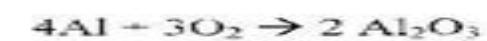
The number of moles of substances, when compared and reduced to the lowest terms, yields a mole ratio, as shown in the sample below. Indicate the mole ratio for each of the following equations:



2: 1: 2



1: 1: 1



4: 3: 2



1: 2: 1 & 1



2: 1: 2



2: 1: 1

Convert the following masses to moles by using the information on your periodic table.

a) 34.0 g NH₃ FW = 17.03/mol

2.00 mol

$$34.0\text{g} \left(\frac{1\text{mol}}{17.03} \right) = 2.00\text{ mol}$$

b) 51.4 g Na₂O FW = 62.09/mol

0.829 mol

$$51.4\text{g} \left(\frac{1\text{mol}}{62.09} \right) = 0.829\text{ mol}$$

c) 4.50 g H₂SO₄ FW = 98.19/mol

0.0459 mol

$$4.50\text{g} \left(\frac{1\text{mol}}{98.19} \right) = 0.0459\text{ mol}$$

d) 229 g Ca(NO₃)₂ FW = 164.19/mol

1.40 mol

$$229\text{g} \left(\frac{1\text{mol}}{164.19} \right) = 1.40\text{ mol}$$

e) 79.5 g NH₄OH FW = 35.09/mol

2.27 mol

$$79.5\text{g} \left(\frac{1\text{mol}}{35.09} \right) = 2.27\text{ mol}$$

Convert the following to grams.

a) 17.1 moles H₂O FW = 18.09/mol

308 g

$$17.1\text{ mol} \times \left(\frac{18.09}{1\text{mol}} \right) = 308\text{ g}$$

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Camilla Rothe

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1. Understanding the eBook Skill Practice 1 Number Practice Answers Chemistry
 - The Rise of Digital Reading Skill Practice 1 Number Practice Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 1 Number Practice Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 1 Number Practice Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 1 Number Practice Answers Chemistry
 - Personalized Recommendations
 - Skill Practice 1 Number Practice Answers Chemistry User Reviews and Ratings
 - Skill Practice 1 Number Practice Answers Chemistry and Bestseller Lists
5. Accessing Skill Practice 1 Number Practice Answers Chemistry Free and Paid eBooks
 - Skill Practice 1 Number Practice Answers Chemistry Public Domain eBooks
 - Skill Practice 1 Number Practice Answers Chemistry eBook Subscription Services
 - Skill Practice 1 Number Practice Answers Chemistry Budget-Friendly Options

6. Navigating Skill Practice 1 Number Practice Answers Chemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - Skill Practice 1 Number Practice Answers Chemistry Compatibility with Devices
 - Skill Practice 1 Number Practice Answers Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 1 Number Practice Answers Chemistry
 - Highlighting and Note-Taking Skill Practice 1 Number Practice Answers Chemistry
 - Interactive Elements Skill Practice 1 Number Practice Answers Chemistry
8. Staying Engaged with Skill Practice 1 Number Practice Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 1 Number Practice Answers Chemistry
9. Balancing eBooks and Physical Books Skill Practice 1 Number Practice Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 1 Number Practice Answers Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Skill Practice 1 Number Practice Answers Chemistry
 - Setting Reading Goals Skill Practice 1 Number Practice Answers Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Practice 1 Number Practice Answers Chemistry
 - Fact-Checking eBook Content of Skill Practice 1 Number Practice Answers Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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