

Chapter 10 Study Guide - KEY

1. What is the molarity of a 2.05 L solution that is made by dissolving 27.5 g of sodium chloride in water?

$$\frac{27.5 \text{ g NaCl}}{58.5 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.470 \text{ mol}$$

$$M = \frac{\text{mol solute}}{\text{L sol'n}} = \frac{0.470 \text{ mol}}{2.05 \text{ L}} = \underline{0.229 M}$$

2. What is the normality of a 6.0 M solution of phosphoric acid?

$$N = (M)(\text{total pos. ox}) = (6.0 M)(3) = \underline{18 N}$$

3. What is the molality of a solution that is made by dissolving 98.7 g of calcium chloride in 755 g of water?

$$\frac{98.7 \text{ g CaCl}_2}{111.1 \text{ g}} \times \frac{1 \text{ mol}}{1} = 0.888 \text{ mol}$$

$$m = \frac{\text{mol solute}}{\text{kg solvent}} = \frac{0.888 \text{ mol}}{0.755 \text{ kg}} = \underline{1.18 m}$$

4. How many grams of KI are needed to prepare 325 mL of a 1.33 M solution?

$$1.33 M = \frac{n \text{ mol}}{0.325 \text{ L}}$$

$$n = 0.432 \text{ mol KI}$$

$$\text{mass} = \frac{0.432 \text{ mol KI}}{1 \text{ mol}} \times \frac{166.1 \text{ g}}{1} = \underline{71.8 \text{ g KI}}$$

5. How many grams of NaCl are needed to prepare a 4.15 m solution, using 176 g of water?

$$4.15 m = \frac{n \text{ mol}}{0.176 \text{ kg}}$$

$$n = 0.7304 \text{ mol NaCl}$$

$$\text{mass} = \frac{0.7304 \text{ mol NaCl}}{1 \text{ mol}} \times \frac{58.5 \text{ g}}{1} = \underline{42.7 \text{ g NaCl}}$$

6. Ethyl iodide ($\text{C}_2\text{H}_5\text{I}$) boils at 72.5°C and has a density of 1.933 g/mL .

- a. A solution prepared by dissolving 0.300 mol of a nonelectrolyte in 750.0 mL of ethyl iodide boils at 73.5°C . What is the boiling point constant (K_b) for ethyl iodide?

$$\text{C}_2\text{H}_5\text{I} = 156.0 \text{ g/mol}$$

$$\text{mass solvent} = \frac{750.0 \text{ mL}}{1 \text{ mL}} \times \frac{1.933 \text{ g}}{1} = 1449.75 \text{ g solvent}$$

$$m = \frac{0.300 \text{ mol}}{1.44975 \text{ kg}} = 0.207 m$$

$$\Delta T_b = K_b \cdot m$$

$$1.0^\circ\text{C} = (K_b)(0.207 m)$$

$$K_b = \underline{4.83^\circ\text{C/m}}$$

Study Guide Solution Concentration

John A. Olmsted, Gregory M. Williams



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Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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Study Guide and Glossary to Accompany Chemical Principles Daniel L. Reger, Edward E. Mercer, Robert S. Boikess, 1978

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