

NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

1. Write the symbol for uranium—238 in isotope format (with top and bottom #s).
2. What does the top # represent?
3. What does the bottom # represent?
4. What particles are found in the nucleus?
5. Why does the strong nuclear force exist?
6. What is the electrostatic force?
7. What particles does the electrostatic force deal with?
8. Define nuclear reactions.
9. Define nuclear fission and fusion.
10. Nuclear reactions usually produce _____ amounts of energy, and isotopes become _____ stable. The _____ number and the _____ number may also change.
11. What is transmutation?
12. Write the beta decay of radium—226.
13. Write the alpha decay of radium—226.
14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
15. Define nuclear radiation.
16. List the 3 radioactive particles in order of INCREASING penetrating power.
17. What is radioactive decay?
18. Alpha particles are basically _____ nuclei, and beta particles are basically high energy _____.
19. What is half life?
20. What is radioactive dating?
21. What is cobalt—60 used for?
22. Does fission or fusion produce nuclei with heavier mass than the original mass?
23. Name 3 differences between chemical and nuclear reactions.
24. What is carbon—13 radioactive?
25. Elements past atomic # _____ are radioactive.
26. If 1/16 of a sample of a radioactive isotope is left after nuclear decay, how long did this take if the $\frac{1}{2}$ life is 4 days?
27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sample is left after 12 days?

Study Guide Nuclear Chemistry

NARAYAN CHANDER

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Chemical Principles Peter Atkins,Loretta Jones,2010 This text is designed for a rigorous course in introductory chemistry Its central theme is to challenge students to think and question while providing a sound foundation in the principles of chemistry

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