

NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

1. Write the symbol for uranium—238 in isotope format (with top and bottom #s).
2. What does the top # represent?
3. What does the bottom # represent?
4. What particles are found in the nucleus?
5. Why does the strong nuclear force exist?
6. What is the electrostatic force?
7. What particles does the electrostatic force deal with?
8. Define nuclear reactions.
9. Define nuclear fission and fusion.
10. Nuclear reactions usually produce _____ amounts of energy, and isotopes become _____ stable. The _____ number and the _____ number may also change.
11. What is transmutation?
12. Write the beta decay of radium—226.
13. Write the alpha decay of radium—226.
14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
15. Define nuclear radiation.
16. List the 3 radioactive particles in order of INCREASING penetrating power.
17. What is radioactive decay?
18. Alpha particles are basically _____ nuclei, and beta particles are basically high energy _____.
19. What is half life?
20. What is radioactive dating?
21. What is cobalt—60 used for?
22. Does fission or fusion produce nuclei with heavier mass than the original mass?
23. Name 3 differences between chemical and nuclear reactions.
24. What is carbon—13 radioactive?
25. Elements past atomic # _____ are radioactive.
26. If $1/16$ of a sample of a radioactive isotope is left after nuclear decay, how long did this take if the $\frac{1}{2}$ life is 4 days?
27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sample is left after 12 days?

Study Guide Nuclear Chemistry

NARAYAN CHANGDER



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Chemical Principles Peter Atkins, Loretta Jones, 2010 This text is designed for a rigorous course in introductory chemistry Its central theme is to challenge students to think and question while providing a sound foundation in the principles of chemistry

Essentials of Nuclear Chemistry Hari Jeevan Arnika, 1987-04-15 A thorough introduction to the essential topics of nuclear chemistry With clarity and illustrative examples it covers nuclear structure and stability types of radioactivity and nuclear reactions and the processes of nuclear fission and fusion This edition offers clearer and more up to date coverage of the subject and incorporates entirely new material as well New to this edition a detailed account of nuclear magnetic resonance coverage of the differences and limitations of the Gamov Teller and Fermi selection rules and examples of the earliest nuclear reactions in the cosmos Special attention is paid to the study of magnetic moments of elementary particles and nuclei Features numerical examples with answers and a unique and helpful inclusion of historically important and interesting events

College Chemistry Steven Boone, Drew H. Wolfe, 2011-11-01 The Collins College Outline for College Chemistry is a comprehensive guide to the fundamental concepts behind chemical reactions bonding equilibria and thermodynamics with topics ranging from simple chemical measurements and the basics of atoms and molecules to entropy electrochemistry and nuclear chemistry Fully revised and updated by Dr Steven Boone College Chemistry includes practical test yourself sections with answers and complete explanations at the end of each chapter Also included are essential vocabulary definitions and sample exercises as well as detailed images charts and diagrams The Collins College Outlines are a completely revised in depth series of study guides for all areas of study including the Humanities Social Sciences Mathematics Science Language History and Business Featuring the most up to date information each book is written by a seasoned professor in the field and focuses on a simplified and general overview of the subject for college students and where appropriate Advanced Placement students Each Collins College Outline is fully integrated with the major curriculum for its subject and is a perfect supplement for any standard textbook

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