

## Study Guide Unit 9

1. KNOW the following vocabulary words (this means you NEED to be able to DEFINE them ALL):

- a. Solution- m. Solute-
- b. Homogeneous- n. Solvent-
- c. Heterogeneous- o. Acid-
- d. Soluble- p. Base-
- e. Insoluble- q. Titration-
- f. Miscible- r. Analyte-
- g. Immiscible- s. Titrant-
- h. Solubility- t. Monoprotic acid-
- i. Alloy- u. Diprotic Acid
- j. Saturated- v. Triprotic Acid
- k. Unsaturated-
- l. Supersaturated-

2. How is the solubility of a gas affected by temp. and pressure?

Pressure up = solubility up

Temp up = solubility down

3. Under what conditions are salts most soluble?

High temperatures

4. Use the solubility curve to the right to answer the following.

a. What mass of  $\text{KNO}_3$  will make a saturated solution at  $60^\circ\text{C}$ ? **100g**

b. Which of the compounds shown on the solubility curve decrease solubility as temperature increases?

$\text{NH}_3$ ,  $\text{Ca}_2(\text{SO}_4)_3$

c. What type of solution do I have if I **add** 110 g of  $\text{NaNO}_3$  in 100g of  $20^\circ\text{C}$  water?

**Saturated**

d. What mass of  $\text{KCl}$  can go into solution at  $80^\circ\text{C}$ ?

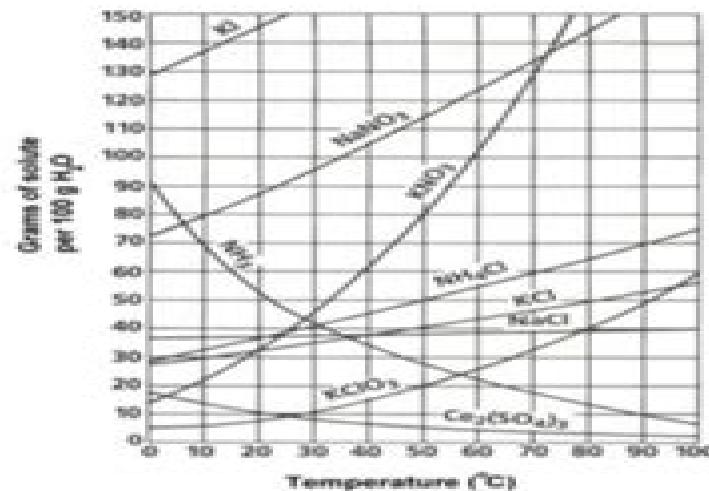
**50 g**

e. What is the least soluble compound at  $20^\circ\text{C}$ ?

**$\text{KClO}_3$**

f. Excluding  $\text{KI}$ , what is the most soluble compound at  $80^\circ\text{C}$ ?

**$\text{KNO}_3$**



# Study Guide Solutions Answer Key

**Carole Reid Clyne, Barbara Trenholm**

## **Study Guide Solutions Answer Key:**

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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**Study Guide for Macroeconomics 2e** Timothy D. Tregarthen, John Brock, Dale Deboer, 2000-01-15 Accounting Principles, Parts 1 and 2

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