



## Unit: Stoichiometry Answer Key

### "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide ( $K_2O$ ) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol K} \times \frac{2 \text{ mol } K_2\text{O}}{4 \text{ mol K}} = 0.760 \text{ mol } K_2\text{O}$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol Cu}_2\text{S} \times \frac{2 \text{ mol Cu}}{1 \text{ mol Cu}_2\text{S}} = 0.50 \text{ mol Cu}$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol Na}_2\text{SO}_4 \times \frac{2 \text{ mol Na}}{1 \text{ mol Na}_2\text{SO}_4} = 7.50 \text{ mol Na}$$

3. How many moles of calcium oxide will be produced when 1.6 moles of iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol Fe}_2\text{O}_3 \times \frac{3 \text{ mol CaO}}{1 \text{ mol Fe}_2\text{O}_3} = 4.8 \text{ mol CaO}$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: magnesium chlorate  $\rightarrow$  magnesium chloride + oxygen

Formula equation:  
(must be balanced!)



$$5.12 \text{ mol Mg}(\text{ClO}_3)_2 \times \frac{3 \text{ mol O}_2}{1 \text{ mol Mg}(\text{ClO}_3)_2} = 15.4 \text{ mol O}_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: chromium + lead (IV) nitrate  $\rightarrow$  chromium (III) nitrate + lead

Formula equation:  
(must be balanced!)



# Study Guide The Mole Answer Key

**William H. Myers, William T. Mooney**

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