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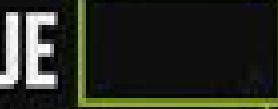
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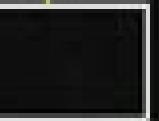
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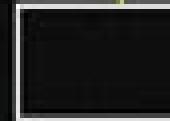
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**Christian Drosten**

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dreams and expectations of young men who for better or worse believe in America's pastime as their gateway out of poverty

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**Current Biography Yearbook 1986**,    **Premises and Conclusions** Robert E. Rodes (Jr.), Howard Pospesel, 1997  
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formulae and equations e g in one mole of carbon c the number of atoms is the same as the number of molecules in one mole of carbon dioxide co 2 linking the mole and the atomic mass

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*the mole 4 2 1 cie igcse chemistry revision notes 2022* - Jul 14 2023

web for practical purposes a rounded version of the constant can be used in exams  $6.02 \times 10^{23}$  mol 1 the reason we have mol 1 as the unit is because this is the number of entities per mole of substance for example one mole of sodium na contains  $6.02 \times 10^{23}$  atoms of sodium one mole of hydrogen h 2 contains  $6.02 \times 10^{23}$  molecules of hydrogen