

1. Determine the mass and the weight of the air contained in a room whose dimensions are  $6\text{ m} \times 6\text{ m} \times 3\text{ m}$ . Assume the density of the air is  $1.16 \text{ kg/m}^3$ .

2. Determine the mass and the weight of the air contained in a room whose dimensions are  $6\text{ m} \times 6\text{ m} \times 3\text{ m}$ . Assume the density of the air is  $1.16 \text{ kg/m}^3$ .

Given :

$$\rho = 1.16 \text{ kg/m}^3 \quad \rightarrow \text{Density of air}$$

$$g = 9.81 \text{ m/s}^2 \quad \rightarrow \text{Assume gravity acceleration}$$

$$6\text{ m} \times 6\text{ m} \times 3\text{ m} \quad \rightarrow \text{Dimensions of room}$$

Solution :

Volume of room :

$$V = 6\text{ m} \times 6\text{ m} \times 3\text{ m} = 216 \text{ m}^3$$

Mass is defined by volume of room & air :

$$m = \rho \cdot V = (1.16 \frac{\text{kg}}{\text{m}^3})(216 \text{ m}^3) = \boxed{334.08 \text{ kg}}$$

Weight of object :

$$G = m \cdot g = (334.08 \text{ kg})(9.807 \frac{\text{m}}{\text{s}^2}) \left( \frac{1 \text{ N}}{1 \text{ kg} \cdot \frac{\text{m}}{\text{s}^2}} \right) = \boxed{3277.73 \text{ N}}$$

Therefore :

$$m = 334.08 \text{ kg}$$

$$G = 3277.73 \text{ N}$$

Answered by D. Quillen ~~10/10/2023~~

# Thermodynamics Problem And Solution

**Kathleen E. Murphy**

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Robert Balmer,2011-01-25 Modern Engineering Thermodynamics is designed for use in a standard two semester engineering thermodynamics course sequence The first half of the text contains material suitable for a basic Thermodynamics course taken by engineers from all majors The second half of the text is suitable for an Applied Thermodynamics course in mechanical engineering programs The text has numerous features that are unique among engineering textbooks including historical vignettes critical thinking boxes and case studies All are designed to bring real engineering applications into a subject that can be somewhat abstract and mathematical Over 200 worked examples and more than 1 300 end of chapter problems provide opportunities to practice solving problems related to concepts in the text Provides the reader with clear presentations of the fundamental principles of basic and applied engineering thermodynamics Helps students develop engineering problem solving skills through the use of structured problem solving techniques Introduces the Second Law of Thermodynamics through a basic entropy concept providing students a more intuitive understanding of this key course topic Covers Property Values before the First Law of Thermodynamics to ensure students have a firm understanding of property data before using them Over 200 worked examples and more than 1 300 end of chapter problems offer students extensive opportunity to practice solving problems Historical Vignettes Critical Thinking boxes and Case Studies throughout the book help relate abstract concepts to actual engineering applications For greater instructor flexibility at exam time thermodynamic tables are provided in a separate accompanying booklet Available online testing and assessment component helps students assess their knowledge of the topics Email [textbooks.elsevier.com](http://textbooks.elsevier.com) for details **Problems in Chemical**

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*Thermodynamics* William Z. Black, James G. Hartley, 1996 The third edition of Thermodynamics provides an easily understandable presentation of classical thermodynamics that builds on the student's background of energy concepts first learned in physics and chemistry. The material is organized in a logical progression from the conservation of mass, the conservation of energy and the second law. The engineering perspective is retained and a variety of familiar examples are used so that the student can appreciate how thermodynamics affects a broad range of subjects. The authors continue to emphasize a systematic approach to problem solving and that approach is used in all example problems in the text. This problem solving method provides not only a reasonable way to approach the task of solving thermodynamics problems but it also will serve the student in other engineering and science disciplines. Each example is worked in detail and particular attention has been given to the proper use of units and unit conversions in the solutions. Detailed explanations accompany the simplifications when the general equations are reduced to the forms that apply to special cases so that the student will gain a better understanding of the conservation principles as well as greater awareness of these powerful analytical tools. Examples address the questions of which form of the conservation laws should be used and why certain assumptions can be applied to simplify the solutions. Believing that second law analysis should play a major role in the analysis of engineering problems, the authors provide extensive coverage of the second law of thermodynamics. The development of the second law is similar to that used for the introduction of the conservation of mass and energy. The results of the second law are carried over into subsequent chapters where they are applied to thermodynamic systems such as power and refrigeration cycles as well as air conditioning processes.

**Problems in Thermodynamics and Statistical Physics** Peter T. Landsberg, 2014-07-16 Well-respected and widely used, this volume presents problems and full solutions related to a wide range of topics in thermodynamics, statistical physics and statistical mechanics. The text is intended for instructors, undergraduates and graduate students of mathematics, physics, chemistry and engineering. Twenty-eight chapters, each prepared by an expert, proceed from simpler to more difficult subjects. Similarly, the early chapters are easier than the later ones, making the book ideal for independent study. Subjects begin with the laws of thermodynamics and statistical theory of information and of

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