

Chemistry 101 Final Exam Study Guide- Spring 2017

Acids and Bases

Acid: substance that provides H⁺ ions in water

Base: substance that provides OH⁻ ions in water

Covalent Bonds

- Review: Ionic bonds are formed by the positive charge of the cation being attracted to the negative charge of the anion
- Covalent bonds: bonds formed by sharing electrons between the atoms (not ions)
 - 1 electron from each atom (2 atoms per bond)
 - Electrons available for bonding are the un-paired electrons in orbitals (orbitals with only 1 electron in them)
- Molecule
 - A group of atoms held together by covalent bonds
- Atoms will bond with enough other atoms to complete the octet
 - But they can only form as many bonds as unpaired valence electrons
 - Hydrogen has 1 valence electron, so it can only form 1 bond, and will not achieve an octet
 - Boron has 3 valence electrons, so can only form 3 bonds, and will not achieve the octet (only 6 electrons)
 - There are always exceptions to the rules
 - Sulfur and phosphorus can extend the octet

Multiple Covalent Bonds between 2 atoms

- If the octet is not met, additional bonds may be formed
 - CO₂ (carbon dioxide)
 - C has 4 valence electrons, so 4 bonds possible
 - O has 6 valence electrons, 2 are unpaired, so 2 bonds are possible

PRACTICE

IONS:

- Sodium
 - Symbol: Na
 - Valence: 1
 - Electron dot: Na⁺
 - Ionic dot: Na⁺ (loses its electron to get a 1+ charge)
- Magnesium
 - Symbol: Mg
 - 2 valence E-
 - Loses 2 electrons to get 2+ charge
- Oxygen
 - Symbol: O
 - Valence: 6
 - Will gain 2 electrons for a 2- charge

Cannot have a covalent bond with ions, just atoms

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