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Instructor, UCSD Computer Science Department, and best-selling author

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web the balanced chemical equation for this reaction is $\text{Mg}(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + \text{H}_2(\text{g})$ in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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b $\text{mg}^{2+} \text{ aq } + \text{co}_3^{2-} \text{ aq } \rightarrow 2\text{h}^+ \text{ aq } + 2\text{no}_3^- \text{ aq }$ mg write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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